

Ch. 12 practice 106, 108, 110, 112

Reading =

Ch. 13.1 - 13.4 and 13.6 - 13.7 and 13.9 - 13.12

Hw #6 due Oct. 16

PE increases as  $Z_{\text{eff}}$  increases

becomes  
more negative  
→ IE is increasing

Exceptions to Trends =

$IE_{\text{Mg}} > IE_{\text{Al}}$  → P<sup>1</sup> electron is shielded by the 2s electrons

$IE_{\text{N}} > IE_{\text{O}}$  →

O =  $\uparrow\downarrow$   $\uparrow$   $\uparrow$   
N =  $\uparrow$   $\uparrow$   $\uparrow$   $\uparrow$

spin pairing energy >  $Z_{\text{eff}}$  in this case

\* spin repulsion occurs any time electrons are paired up

He is hardest to ionize:  $Z_{\text{eff}}$  is 23.72

$$E \propto \frac{Z_{\text{eff}}}{n^2}$$

valence electrons don't shield well as  $Z_{\text{eff}}$  increases

Group 1A



(i) 2nd series  $n=3$   
 $3s \rightarrow \text{Na}$



(ii)  $4s \rightarrow \text{K}$



(iii)  $4s \rightarrow \text{K}$



(iv)  $2s \rightarrow \text{Li}$

increasing  
atomic radius

Li  
↓  
Na  
↓  
K  
↓  
Rb

decreasing  
ionization energy

Rb  
↓  
K  
↓  
Na  
↓  
Li

# Chem 101 Upenn

**Karen R. Diaz**



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