

Questions and Problems

Q1 How many mL of a 0.10 M NaOH solution are needed to neutralize 15 mL of a 0.20 M H_3PO_4 solution?



B. Titration of an Antacid

	Antacid 1	Antacid 2
1. Brand of antacid	Tums	A-2
Bases(s) in antacid	CaCO_3	NaHCO_3
2. Mass of flask	178.64 g	178.55 g
3. Mass of flask and antacid	179.11 g	178.93 g
4. Molarity of HCl solution	0.10 M	0.10 M
5. Total volume (mL) of HCl solution added	50.0 mL	50.0 mL
6. Molarity of NaOH solution	0.10 M	0.10 M
7. Initial volume (mL) of NaOH solution	0.54 mL	0.58 mL
8. Final volume (mL) of NaOH solution	14.32 mL	23.70 mL
9. Mass of antacid	0.47 g	0.38 g
10. Volume of NaOH solution used		
11. Volume of excess HCl solution		
12. Volume of HCl solution neutralized by antacid		
13. $\frac{\text{mL stomach acid}}{1 \text{ g antacid}}$		

14. Write the neutralization equations that take place in the stomach with the base(s) present in the antacid product.

- _____
- _____

Q2 How many grams of $\text{Mg}(\text{OH})_2$ will be needed to neutralize 25 mL of stomach acid if stomach acid is 0.10 M HCl?

Chem 122 Lab Manual Answers General Organic

**Detroit (Mich.) College of the city of
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