

Ideal Gas Law, $PV = nRT$

Name _____

Instructor _____ Section _____ Partner _____

(Unless otherwise noted, 2 pts. for each blank)

	Data, Trial 1	Data, Trial 2
Molarity of HCl (M) (1pt. each):	_____	_____
Volume of HCl (mL):	<u>19.9 mL</u>	<u>20.0 mL</u>
Mass of Mg (gram):	<u>0.044 g</u>	<u>0.046 g</u>
*Volume of gas before placing in Equalization Chamber (mL) (1 pt.)	<u>45.1 mL</u> (*This value NOT used in calculations below)	<u>46.1 mL</u>
Volume of gas after placing in Equalization Chamber (mL):	<u>44.9 mL</u>	<u>46.0 mL</u>
Barometric Pressure (mm Hg):	<u>738.2 mmHg</u>	<u>738.2 mmHg</u>
Temperature (°C):	<u>22.0 °C</u>	<u>22.0 °C</u>

Write the balanced overall equation for the reaction of magnesium and HCl to form hydrogen gas and the magnesium salt (2 pts):

Mole of Mg reacted	_____	_____
Mole of H ₂ formed (based on the stoichiometric equation for the reaction)	_____	_____
Vapor Pressure of water from curve in mmHg:	_____	_____
Pressure of H ₂ from Dalton's Law in mmHg:	_____	_____
Pressure of H ₂ in atm	_____	_____
Volume of H ₂ in liter (Using volume measured in Equalization Chamber)	_____	_____
Temperature in Kelvin	_____	_____
Calculation of R (6 pts. each)	_____	_____
Percent error	_____	_____
Average R =	_____	_____

Chem 1211 Lab Manual Answers

Rosina Ehmman



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