



ANSWER KEY

How many atoms of **oxygen (O)** are in the **chemical formula** below?

- A 4
- B 5
- C 7
- D 12



(d)

The diagram shown below is a **structural formula** for the gas **propane**. What would be the correct **chemical formula** for propane?

- A C_3H_8
- B CH_4
- C $8\text{H}_3\text{C}$
- D CH



(a)

Carbon (C) and hydrogen (H) create **covalent bonds** to share **electrons** between them. What is the total number of electrons being shared between **carbon and hydrogen** in the molecule of propane shown below?

- A 8
- B 11
- C 16
- D 32



(c)

The **reason** why the bonds between carbon (C) and hydrogen (H) are **covalent** is _____

- A C and H do not have electrons to give away
- B C and H are both nonmetals
- C C and H are both metalloids
- D C and H can bond only by sharing electrons



(d)

The diagram below represents the **structural formula** for several **water molecules**. How many **molecules** are represented?

- A 3
- B 5
- C 10
- D 15



(b)

Explain the **arrangement** of the molecules in the diagram of water below.

- A It is a lucky arrangement
- B water molecules are sticky
- C water molecules have polarity
- D water molecules are magnetic



(c)

Water molecules have some **polarity** in their structure. Using the diagram below, notice how the ends of H atoms (**positive**) and O atoms (**negative**) meet. This property is the **reason** that _____

- A water forms droplets easily
- B water boils easily
- C ice melts easily
- D water has color



(a)

Since H atoms are slightly **positive** and the O atoms are slightly **negative**, molecules of water have **polarity**. This makes water a **good** _____

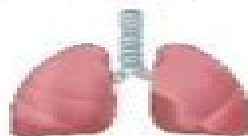
- A solute
- B refrigerant
- C solvent
- D magnet



(c)

The chemical formula for the **oxygen** that we breathe is _____

- A O
- B O_2
- C O_3
- D 2O



(b)

How many **molecules of oxygen** are represented in the chemical formula below?

- A 2
- B 4
- C 6
- D 18



(b)

Chemistry Study Guide Chemical Bonding

**Barron's Educational Series, Gabrielle
I. Edwards, Cynthia Pfirrmann**



Chemistry Study Guide Chemical Bonding:

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