

Pg. 1 Basic Skills; Math Review • Pg. 2 Statistics, Atomic Data, Chemical Formula Calculations • Pg. 3 Stoichiometry  
Pg. 4 Gas Law Calculations, Solids and Liquids • Pg. 5 Thermodynamics, Acid-Base Calculations • Pg. 6 Equilibrium Calculations, Kinetics

Your success in chemistry depends on your ability to solve numerical problems.

## BASIC SKILLS

### CALCULATOR SURVIVAL

- Become familiar with your calculator.
- Before the exam, make sure you can multiply, divide, add, subtract and use all needed functions.
- Calculators never make mistakes; they take your input (rounded off, accidentally and give an answer.
- Look at the answer; does it make sense?
- Do a quick estimate to check your work.
- Sample:**  $4.34 \times 7.88 \div 1.55 \times 9.81$  is roughly  $4 \times 8 \div 1.5 \times 10$  or 3.2; the actual answer, 3.24

### HOW TO DO WORD PROBLEMS

- Read and evaluate the question before you start plugging numbers into the calculation.
- Identify the variables, constants and equations.
- Write out units of the variables and constants.
- Work out the unit before the number.
- You may have extra information, or you may need to obtain constants from your text.

### HOW TO WORK WITH UNITS

All numerical data has units. In chemistry we use metric "SI" units.

**Pitfall:** If the unit is wrong, the answer is wrong!

1. **Unit prefixes:** denote powers of  $10^3$ 
  - tera T  $10^{12}$
  - kilo k  $10^3$
  - milli m  $10^{-3}$
  - micro  $\mu$   $10^{-6}$
  - pico p  $10^{-12}$

### 2. Fundamental Units

- Mass: kilogram (kg)
- Length: meter (m)
- Temperature: kelvin (K)
- Time: second (s)
- Amount of a substance: mole
- Electric charge: coulomb (C)

### 3. Derived Units

- Area: length squared,  $m^2$ ; 1 liter (L) = 1  $dm^3$
- Volume: length cubed,  $m^3$ ; 1 liter (L) = 1  $dm^3$
- Density: mass/volume; common unit  $kg/m^3$
- Speed: distance/time; common unit  $m/s$
- Electric current: ampere (A) = 1 C/s
- Force: Newton (N) = 1 kg  $m/s^2$
- Energy: joule (J) = 1 kg  $m^2/s^2$
- Pressure: Pascal (Pa) = 1  $kg/m \cdot s^2$

### 4. Fundamental Constants

|                            | in units (symbol)                                      | in kg                                                  |
|----------------------------|--------------------------------------------------------|--------------------------------------------------------|
| Avogadro's number, $N_A$   | $6.022 \times 10^{23}$                                 | $6.022 \times 10^{23}$                                 |
| Ideal gas constant, R      | $8.314 \text{ J/K} \cdot \text{mol}$                   | $8.314 \text{ J/K} \cdot \text{mol}$                   |
| R, for gas calculations    | $0.0821 \text{ L} \cdot \text{atm/K} \cdot \text{mol}$ | $0.0821 \text{ L} \cdot \text{atm/K} \cdot \text{mol}$ |
| R, for energy calculations | $8.314 \text{ J/K} \cdot \text{mol}$                   | $8.314 \text{ J/K} \cdot \text{mol}$                   |
| Faraday's Constant, F      | $96,485 \text{ C/mol}$                                 | $96,485 \text{ C/mol}$                                 |
| Planck's Constant, h       | $6.626 \times 10^{-34} \text{ J} \cdot \text{s}$       | $6.626 \times 10^{-34} \text{ J} \cdot \text{s}$       |
| Speed of Light, c          | $2.9979 \times 10^8 \text{ m/s}$                       | $2.9979 \times 10^8 \text{ m/s}$                       |

## BASIC SKILLS cont.

### HOW TO CONVERT DATA

The unit & numerical value are changed using a conversion factor or equation.

### HOW TO USE EQUATIONS FOR DATA CONVERSION

- Sample:** Convert a temperature of 43°C to °F and Kelvin
- Given:  $K = 1C + 273.15$
- Calculate: Temperature in K =  $43^\circ\text{C} + 273.15 = 316.15 \text{ K}$
- Given:  $F = 9/5 \cdot C + 32$
- Calculate: Temperature in °F =  $9/5 \times 43^\circ\text{C} + 32 = 111 + 32 = 143^\circ\text{F}$



### HOW TO USE A CONVERSION FACTOR

- Sample:** "1 hour = 60 minutes" gives two conversion factors.
- 1. Multiply by "1 hour/60 minutes" to convert minutes to hours.
- 2. Multiply by "60 minutes/1 hour" to convert hours to minutes.
- Sample:** Determine the number of hours in 43 minutes.
- Given: Conversion factor is "1 hour/60 minutes"
- Calculate: Time = 43 minutes  $\times$  (1 hour/60 minutes) = 0.72 hour (rounded correctly)

### Common conversions:

|                                          |                                                |              |
|------------------------------------------|------------------------------------------------|--------------|
| 1 caloric = 4.184 J                      | 1 kg = 2.2 lb                                  | 1 m = 1.1 yd |
| 1 gal = 3.785 L                          | 1 angstrom (Å) = $1 \times 10^{-10} \text{ m}$ |              |
| 1 atm = 1.01325 $\times 10^5 \text{ Pa}$ | 1 mm = 10 mm Hg                                |              |

**Pitfall:** Equations and conversion factors have units. Be sure of inverting conversion factors.

## MATH REVIEW

### HOW TO WORK WITH ALGEBRAIC EQUATIONS

Give equal treatment to each side.

- **Add or subtract:**
- Sample:** Given:  $x + y = 5$ , then  $4 + x = 4 + 5$
- **Multiply or divide:**
- Sample:** Given:  $x = y$ , then  $4x = 4y$  and  $x/2 = y/2$
- Given:  $ax + b = 3y + 2z$ , then  $4ax + 4b = 12y + 8z$ ,  $x = 2$
- Pitfall:** Dividing by zero is not allowed.

### HOW TO WORK WITH SCIENTIFIC NOTATION

The exponent gives the power of 10.

- Sample:**  $0.00045 = 4.5 \times 10^{-4}$ ;  $0.45 = 4.5 \times 10^{-1}$
- Sample chemical applications:** Molecular diameters are  $10^{-10} \text{ m}$ ; 1 liter of water contains about  $1 \times 10^{25}$  atoms.

## MATH REVIEW cont.

### HOW TO USE LOGARITHMS & EXPONENTS

- **Common logarithm,  $\log_{10}$  base "10"**
- Denotes number or function in powers of 10.
- Sample:** Given:  $y = 10^5$ , then  $\log_{10} y = 5$ .
- **Natural logarithm,  $\ln$  base "e"**
- Denotes number or function in powers of "e".
- Sample:** Given:  $x = e^3$ , then  $\ln x = 3$ .
- **Products:**  $\log(xy) = \log x + \log y$
- **Powers:**  $\log(x^y) = y \log x$
- **Multiplication:** Add exponents from each term.
- Sample:**  $10^3 \times 10^2 = 10^{3+2} = 10^5$
- **Division:** Subtract denominator exponents from numerator exponents.
- Sample:**  $10^5/10^2 = 10^{5-2} = 10^3$
- **Square root:**  $\sqrt{x} = x^{1/2}$
- **Inverse:**  $1/x = x^{-1}$

**Pitfall:** Your calculator has separate keys for  $\ln$  (base e),  $\log$  (base 10),  $10^x$  and  $x^y$ .  
Sample chemical applications: pH of acid and base.

### HOW TO CALCULATE ROOTS OF A POLYNOMIAL

An equation of the form  $ax^2 + bx + c = 0$ , has 2 solutions or roots, given by the quadratic formula:

$$x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$$

- Sample:** Determine the roots for the equation:  $3x^2 + 4x + 1 = 0$
- Given:  $a = 3$ ,  $b = 4$ ,  $c = 1$
- Calculate: Root 1 =  $\frac{-4 + \sqrt{4^2 - 4 \times 3 \times 1}}{2 \times 3} = \frac{-4 + \sqrt{16 - 12}}{6} = \frac{-4 + 2}{6} = -\frac{2}{6} = -\frac{1}{3}$
- Calculate: Root 2 =  $\frac{-4 - \sqrt{16 - 12}}{6} = \frac{-4 - 2}{6} = -\frac{6}{6} = -1$

**Pitfall:** Always round-off correctly. Substitute the roots into the equation to verify results.  
Sample chemical applications: Weak acids, weak bases, buffers, chemical equilibrium

### HOW TO DETERMINE THE EQUATION OF A LINE

**LINEAR EQUATION:  $y = mx + b$**

- $m$ : slope of the line;  $m = \Delta y/\Delta x$
- $b$ : y intercept, the line crosses the y-axis at  $(0, b)$  or  $y = -mx$ .
- Sample:** Determine the equation of a line using (x,y) data.
- Given:  $(-2, -4)$ ,  $(0, -2)$ ,  $(2, 0)$
- Calculate: Slope =  $m = \Delta y/\Delta x$
- $= (12 - (-20)) / (2 - (-2)) = 32 / 4 = 8$
- Calculate: y-intercept =  $b = y_1 - mx_1$
- $= (-4) - (8 \times (-2)) = -4 + 16 = 12$

The equation of the line is  $y = 8x + 12$ .  
Sample chemical applications: Gas calculations; Beer's Law; analyzing reaction-rate data



# Chemistry Study Guide Answers Concepts And

**Kathleen Armour**



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