Topic 4: Stoichiometry

1) Chemical Formula

4) Moles (3 laws)

2) Chemical Equations

- 5) Empirical formula & Molecular Formula
- 3) Relative Molecular/Formula mass (Mr)

1) Chemical Formula

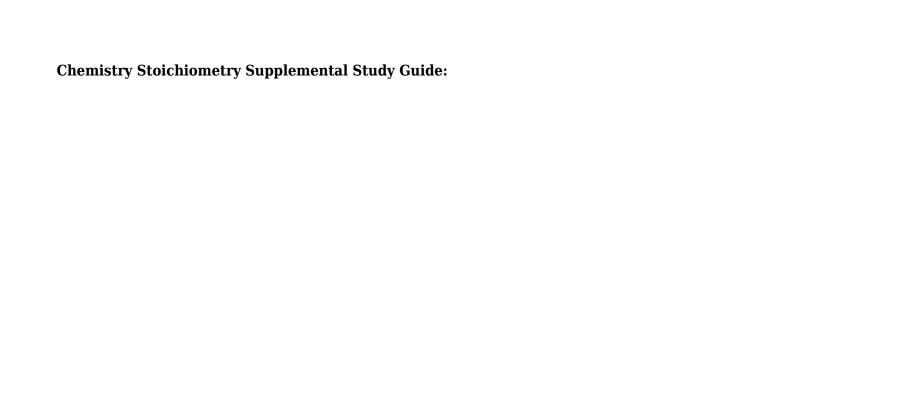
Valency or combining power:

- The valency of an element is the number of electrons lost, gained or shared to from compounds. The valency of an element depends on the number of valence electrons.
- In covalent molecules; the valency gives the number of covalent bonds which the atoms can form.
- In ionic compounds; the valency gives the charge on the ions of the element.
- The following table shows the valencies & charges of common elements used in chemistry:

Common metals	Element	Symbol	Valency	Charge	Common Non Metals	Element	Symbol	Valency	Charge	
	Lithium	Li	1	•		Hydrogen	н	1	+1	
	Sodium	Na	1	**		Chlorine	CI	1	-1	
	Potassium	ĸ	1	+.		Bromine	Br	1	-1	
	1/2 Silver	Ag	1			lodine	- 1	1	-1	
	Calcium	Ca	2	+2		Oxygen	0	2	-2	
	Barium	Ba	2	+2		Sultur	s	2	-2	
	⊀r Lead	Ръ	2	+2		Nitrogen	N	3	-3	
	Magnesium	Mg	2	+2		Phosphonus	P	3	-3	
	Moreury	Hg	2	+2		Carbon	C	4	- 4	
	Zinc	Zn	2	+2		Sillicon	Si	4	- 4	
	Aluminium	AI	3	+3	N.B:					
	Iron II	Fe	2	+2	١.	Element + Oxygen → Oxide Element + Sulfur → Sulphide Element + Nitrogen → Nitride Element + Phosphorus → Phosphide Element + Hallogen → Hallide				
	Iron III	Fe	3	+3						
	Copper I	Cu	1	+1						
	Copper II	Cu	2	+2	Element + Carbon → Carbide Element + Hydrogen → Hydride					

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