

**Chemistry: Stoichiometry: Problem Sheet 2**

**Directions:** Solve each of the following problems. Show your work, including proper units, to earn full credit.

1. Mass percent composition according to the following chemical equation:



(a) How many moles of gas are needed to produce 50 moles of water?

$$x \text{ moles H}_2\text{O}_2 + 50 \text{ moles H}_2\text{O} \left( \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles H}_2\text{O}_2} \right) = 50 \text{ moles H}_2\text{O}$$

(b) How many moles of water are produced by 50 moles of gas in the reaction?

$$x \text{ moles H}_2\text{O}_2 + 50 \text{ moles H}_2\text{O} \left( \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles H}_2\text{O}_2} \right) = 50 \text{ moles H}_2\text{O}$$

(c) How many moles of gas are needed to produce 50 moles of water?

$$x \text{ moles H}_2\text{SO}_4 + 50 \text{ moles H}_2\text{O} \left( \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles H}_2\text{SO}_4} \right) = 50 \text{ moles H}_2\text{O}$$

(d) How many moles of gas are needed to produce 50 moles of water?

$$x \text{ moles H}_2\text{SO}_4 + 50 \text{ moles H}_2\text{O} \left( \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles H}_2\text{SO}_4} \right) = 50 \text{ moles H}_2\text{O}$$

2. Mass percent composition:  $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$   $2\text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$

(a) How many moles of hydrogen peroxide are required to produce 50 moles of oxygen?

$$x \text{ moles H}_2\text{O}_2 + 50 \text{ moles O}_2 \left( \frac{2 \text{ moles H}_2\text{O}_2}{1 \text{ mole O}_2} \right) = 100 \text{ moles H}_2\text{O}_2$$

(b) How many moles of hydrogen peroxide are required to produce 50 moles of oxygen?

$$x \text{ moles H}_2\text{O}_2 + 50 \text{ moles O}_2 \left( \frac{2 \text{ moles H}_2\text{O}_2}{1 \text{ mole O}_2} \right) = 100 \text{ moles H}_2\text{O}_2$$

(c) How many moles of water are produced from 50 moles of oxygen and oxygen?

$$x \text{ moles H}_2\text{O} + 50 \text{ moles O}_2 \left( \frac{2 \text{ moles H}_2\text{O}}{1 \text{ mole O}_2} \right) = 100 \text{ moles H}_2\text{O}$$

3. Calculating the mass of hydrogen gas produced from 5.00 moles of hydrogen peroxide:

$$x \text{ moles H}_2\text{O}_2 + 5.00 \text{ moles O}_2 \left( \frac{2 \text{ moles H}_2\text{O}_2}{1 \text{ mole O}_2} \right) \left( \frac{2 \text{ moles H}_2\text{O}}{2 \text{ moles H}_2\text{O}_2} \right) = 10.00 \text{ moles H}_2\text{O}$$

Answer: (a) 10.00 moles (b) 10.00 moles (c) 10.00 moles (d) 10.00 moles

# Chemistry Stoichiometry Problem Sheet 2 Key

**Joacim Rocklöv**



## Chemistry Stoichiometry Problem Sheet 2 Key:

The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math. The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors. This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB

Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills, calculation methods and the language and basic concepts of chemistry. The book is designed to meet the specific needs of underprepared students. Concepts are presented only as they are needed and developed from the simple to the complex. The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter, energy and measurement, the properties of atoms, description of chemical bonding, study of chemical change and nuclear and organic chemistry. Undergraduate students will find the book as a very valuable academic material.

*Resources in Education*, 1982-04

**Analytical Chemistry: Key to Progress on National Problems** W. Wayne Meinke, John Keenan Taylor, 1972

Essays in Contemporary Chemistry Gerhard Quinkert, 2001 In this volume internationally renowned chemists recount their roles in the progress of chemistry research toward elucidation of biological processes. Beginning with a historical perspective on the development of X-ray crystallography, the reader is regaled with first-hand accounts of research milestones. Included are descriptions of the cutting-edge nuclear magnetic resonance and electron paramagnetic resonance spectroscopic techniques, the dynamics of ultra-fast reactions and the central problem of molecular oxygen activation in biological processes. The roles of chiral auxiliaries in organic synthesis and of photochemistry in natural product synthesis are told and innovations in combinatorial synthesis approaches are described. Contemporary issues in molecular recognition and modulation of molecular function are addressed, concluding with a missive regarding how the frontiers of medical ethics may be breached by molecular manipulations. The contributors, who number among the finest scientists in the world including two Nobel Prize winners, are Peter B. Dervan, Jack D. Dunitz, Christian Griesinger, Jean-Marie Lehn, Thomas F. Prisner, Gerhard Quinkert, Peter G. Schultz, Helmut Schwarz, Dieter Seebach and Ahmed Zewail. Additionally, there is a prologue by Albert Eschenmoser for whom this collection was conceived and an epilogue that contains facsimiles of notes from his landmark lecture, "Synthesis of Co-Enzyme B<sub>12</sub>: A Vehicle for Teaching Organic Synthesis." This book is definitely a must for all who want to read or to read again where we stand in our chemical comprehension of the fascinating relationship between chemical structure and biological processes, how we got here and what the future might hold.

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02

100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

**A Problem-Solving Approach to Aquatic Chemistry** James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands

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**Chemistry** Daryle Hadley Busch, Harrison Shull, Robert T. Conley, 1978 **Research in Education**, 1974 **Multiple Representations in Chemical Education** John K. Gilbert, David Treagust, 2009-02-28 Chemistry seeks to provide qualitative and quantitative explanations for the observed behaviour of elements and their compounds Doing so involves making use of three types of representation the macro the empirical properties of substances the sub micro the natures of the entities giving rise to those properties and the symbolic the number of entities involved in any changes that take place Although understanding this triplet relationship is a key aspect of chemical education there is considerable evidence that students find great difficulty in achieving mastery of the ideas involved In bringing together the work of leading chemistry educators who are researching the triplet relationship at the secondary and university levels the book discusses the learning involved the problems that students encounter and successful approaches to teaching Based on the reported research the editors argue for a coherent model for understanding the triplet relationship in chemical education

**Advanced Placement Examination in Chemistry** Frederick J. Rowe, 1990 New for 1990 this is the only preparation guide available for this exam Includes analyses of past tests review of chemistry topics and samples of all multiple choice and free response problems

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**Mathematics Formative Assessment, Volume 2** Page Keeley, Cheryl Rose

Tobey,2016-12-08 Everything you need to promote mathematical thinking and learning Good math teachers have a robust repertoire of strategies to move students learning forward This new volume from award winning author Page Keeley and mathematics expert Cheryl Rose Tobey helps you improve student outcomes with 50 all new formative assessment classroom techniques FACTS that are embedded throughout a cycle of instruction Descriptions of how the FACTs promote learning and inform teaching including illustrative examples support the inextricable link between instruction and learning Useful across disciplines Keeley and Tobey s purposeful assessment techniques help K 12 math teachers Promote conceptual understanding Link techniques to core ideas and practices Modify instruction for diverse learners Seamlessly embed formative assessment throughout the stages of instruction Focus on learning targets and feedback Instead of a one size fits all approach you can build a bridge between your students initial ideas and correct mathematical thinking with this one of a kind resource

*Chemistry* Greg Curran,2011 Covers all the topics in a typical one year high school chemistry curriculum

**CliffsNotes Chemistry Practice Pack** Charles Henrickson,2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level

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