

# Worksheet #1: Significant Figures & Dimensional Analysis

Directions: For each problem below, write the equation and show your work. Be sure to **box your final answer**.

Part 1: Solve the following problems using scientific notation and rounding to the appropriate value.

a.  $300,235,800 \times 10^{-9}$

c.  $6,000,235,800 \times 10^{-9}$

b.  $450,500 \times 10^{-6}$

d.  $-60,000 \times 10^{-2}$

→ Dig first @ the result is the same as the  $\neq$  in the least precise measurement.

Part 2: Solve the following problems, and show your final answer with the appropriate number of significant figures.

a.  $1.24056 + 75.80 = 77.04$

d.  $43.8 \times 8.0 + 89.32 \div 75 = 411.2$

b.  $10^{18} + 99(34.5 - 20.1) = 1.21$

e.  $(2.88 + 53) \div (23,000 - 0.77) = 9.9 \times 10^{-4}$

c.  $0.8897 \times 2.15 + 0.0021 = 1.9$

→ same # of decimal places as the least precise measurement.

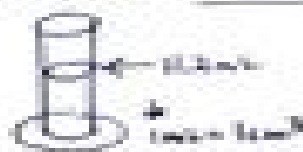
Part 3: For each problem below, show your work. Always use units and box in your final answer.

a. The density of pure silver is  $10.5 \text{ g/cm}^3$  at  $20^\circ\text{C}$ . If  $5.25 \text{ g}$  of pure silver pellets are added to a graduated cylinder containing  $11.2 \text{ mL}$  of water, to what volume level will the water in the cylinder rise?

$d = 10.5 \text{ g/cm}^3$

$m = 5.25 \text{ g}$

$V = ?$



$d = \frac{m}{V}$

$10.5 = \frac{5.25}{x}$

$x = 0.5 \text{ cm}^3$

$11.2 \text{ mL} + 0.5 \text{ mL} = 11.7 \text{ mL}$

b. The density of air at ordinary atmospheric pressure and  $25^\circ\text{C}$  is  $1.19 \text{ g/L}$ . What is the mass, in kilograms, of the air in a room that measures  $12.5 \times 15.5 \times 8.0 \text{ ft}$ ?

$d = 1.19 \text{ g/L}$

$1550 \text{ ft}^3 \cdot \left( \frac{28.32 \text{ L}}{1 \text{ ft}^3} \right) = 43896 \text{ L}$

$52.2 \text{ kg}$

$V = 12.5 \times 15.5 \times 8.0 = 1550 \text{ ft}^3$

$d = \frac{m}{V} \rightarrow 1.19 = \frac{x}{43896} \Rightarrow 52236.24 \text{ g}$

c. An aluminum block has a density of  $2.70 \text{ g/mL}$ . If the mass of the block is  $24.60 \text{ g}$ , find the volume of the substance.

$d = 2.70 \text{ g/mL}$

$m = 24.60 \text{ g}$

$V = ?$

$d = \frac{m}{V} \rightarrow 2.70 = \frac{24.60}{x} = 9.111 \approx 9.11 \text{ mL}$

Part 4: Convert the following measurements to the desired unit.

a.  $0.0003 \text{ cm} = 0.50 \text{ mm}$

b.  $1872 \text{ mg} = 0.001872 \text{ kg}$

c.  $18 \text{ d} = 19 \text{ d}$

d.  $3.4 \times 10.3 \text{ ks} = 390 \text{ s}$

→  $10^3 \text{ (sec)}$   
→  $10^3 \text{ (sec)}$   
→  $10^3 \text{ (sec)}$

# Chemistry States Of Matter Packet Answers Key

**Chris Sunley**



## **Chemistry States Of Matter Packet Answers Key:**

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