

Worksheet #1: Significant Figures & Dimensional Analysis

Directions: For each problem below, write the equation and show your work. Be sure to **box your final answer**.

$$\times / \div$$

Part 1: Solve the following problems using scientific notation and rounding to the appropriate value.

a. $300,234,800 \rightarrow 3.0 \times 10^8$

c. $0.000027100 \rightarrow 2.71 \times 10^{-5}$

b. $450,500 \rightarrow 4.5 \times 10^5$

d. $-60,000 \rightarrow -6.0 \times 10^4$

→ Sig figs in the result is the same as the \pm in the least precise measurement.

Part 2: Solve the following problems, and show your final answer with the appropriate number of significant figures.

a. $1.24056 + 75.80 \rightarrow 77.04$

d. $43.0 \times 8.0 + 89.20 / 75 \rightarrow 411.2$

b. $10^{10} + 99(34.5 - 20) \rightarrow 1.21$

e. $(2.88 + 5) \div (23,000 - 0.71) \rightarrow 9.9 \times 10^{-7}$

c. $0.8097 \times 2.15 + 0.0001 \rightarrow 1.7$

$$+ / -$$

→ same # of decimal places as the least precise measurement.

Part 3: For each problem below, show your work. Always use units and box in your final answer.

1.

a. The density of pure silver is 10.5 g/cm^3 at 20°C . If 5.25 g of pure silver pellets are added to a graduated cylinder containing 11.2 mL of water, to what volume level will the water in the cylinder rise?

$d = 10.5 \text{ g/cm}^3$
 $m = 5.25 \text{ g}$
 $v = ?$



$d = \frac{m}{v} \rightarrow 10.5 = \frac{5.25}{x}$
 $x = 0.5 \text{ cm}^3$

$11.2 \text{ mL} + 0.5 \text{ mL} = 11.7 \text{ mL}$

b. The density of air at ordinary atmospheric pressure and 25°C is 1.19 g/L . What is the mass, in kilograms, of the air in a room that measures $12.5 \times 15.5 \times 8.0 \text{ ft}$?

$d = 1.19 \text{ g/L}$

$1550 \text{ ft}^3 \cdot \left(\frac{28.32 \text{ L}}{1 \text{ ft}^3} \right) = 43896 \text{ L}$

52.2 kg

$28.32 \text{ L} = 1 \text{ ft}^3$
 (Use factor for more info)

$v = 12.5 \times 15.5 \times 8.0 = 1550 \text{ ft}^3$

$d = \frac{m}{v} \rightarrow 1.19 = \frac{x}{43896} \Rightarrow 52236.24 \text{ g} \rightarrow 52.2 \text{ kg}$

c. An aluminum block has a density of 2.70 g/mL . If the mass of the block is 24.60 g , find the volume of the substance.

$d = 2.70 \text{ g/mL}$

$m = 24.60 \text{ g}$

$v = ?$

$d = \frac{m}{v} \rightarrow 2.70 = \frac{24.60}{x} = 9.111 \approx 9.11 \text{ mL}$

Part 4: Convert the following measurements to the desired unit.

a. $0.0001 \text{ cm} = 0.001 \text{ mm}$

b. $1872 \text{ mg} = 0.001872 \text{ kg}$

c. $1 \text{ ft} = 12 \text{ in}$

d. $3.6 \times 10^3 \text{ s} = 390 \text{ min}$

$1 \text{ min} = 60 \text{ sec}$
 $1 \text{ hr} = 60 \text{ min}$

Chemistry States Of Matter Packet Answers Key

Chris Sunley



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