

# CHEMISTRY REVISION GUIDE

## for IGCSE Coordinated Science

This revision guide is designed to help you study for the chemistry part of the IGCSE Coordinated Science course.

The guide contains everything that the syllabus says you need you need to know, and nothing extra.

The material that is only covered in the supplementary part of the course (which can be ignored by core candidates) is highlighted in dashed boxes:



Whilst this guide is intended to help with your revision, it should not be your only revision. It is intended as a starting point but *only a starting point*. You should make sure that you also read your text books and use the internet to supplement your study in conjunction with your syllabus document.

Whilst this guide does contain the entire syllabus, it just has the bare minimum and is not in itself sufficient for those candidates aiming for the highest grades. If that is you, you should make sure you read around a range of sources to get a deeper knowledge and understanding.

Some very useful websites to help you further your understanding include:

- <http://www.docbrown.info/> - whilst not the prettiest site this contains a lot of very useful and nicely explained information.
- <http://www.bbc.co.uk/schools/gcsebitesize/science/> - well presented with many clear diagrams, animations and quizzes. Can occasionally lack depth.
- <http://www.chemguide.co.uk/> - whilst mostly targeted at A-Levels this site contains very detailed information suitable for those looking to deepen their knowledge and hit the highest grades.

Finally, remember revision is not just reading but should be an active process and could involve:

- Making notes
- Condensing class notes
- Drawing Mind-maps
- Practicing past exam questions
- Making flashcards

The golden rule is that what makes you think makes you learn.

*Happy studying, Mr Field.*

### C1: THE ELEMENTS OF CHEMISTRY

Atom: The smallest particle of matter	An atom: 	Some atoms: 
Molecule: A small particle made from more than one atom bonded together	Molecules of an element: 	Molecules of a compound: 
Element: A substance made of only one type of atom	A solid element: 	A gaseous element: 
Compound: A substance made from two or more different elements bonded together	A solid compound: 	A gaseous compound: 
Mixture: A substance made from two or more elements or compounds mixed but not joined	A mixture of compounds and elements: 	

#### CHEMICAL FORMULAS

Formulas tell you the atoms that make up a compound

Eg 1.  $H_2O$  - two H, one O  
Eg 2.  $C_2H_6O$  - two C, six H, one O

Eg 3.  $Mg(OH)_2$  - one Mg, two O, two H

Eg 4.  $CH_3(CH_2)_3$  - three C, 8 H

\* In this case everything in brackets is doubled

#### CHEMICAL EQUATIONS

-Show the reactants you start with and the products you make

-Must contain an arrow ( $\rightarrow$ ) NOT an equals sign (=)

-Must be balanced - same number of atoms on each side

Eg.  $CH_4 + O_2 \rightarrow CO_2 + H_2O$   
This is unbalanced as there are 4 'H' on the left but only 2 'H' on the right. This must be corrected by adding a '2' in front of the ' $H_2O$ ':  
 $CH_4 + O_2 \rightarrow CO_2 + 2H_2O$   
Now the H balances but there 4 O on the right and only 2 on the left. This must be balance by placing a '2' in front of the ' $O_2$ ':  
 $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$   
Now there is 1 'C', 4 'H' and 4 'O' on each side so it balances.

#### MOLES AND MOLAR MASS

-A 'mole' is the same we give to the number  $6.02 \times 10^{23}$  - it is used to talk about the numbers of particles involved in chemical reactions.

-It is the number of atoms such that one mole has the same mass in grams as an atom's atomic mass.

-Eg 1: Carbon has an atomic mass of 12.0, so a mole of carbon has a mass of 12.0g

-Eg 2: Iron has an atomic mass of 55.8, so a mole of iron has a mass of 55.8g

-The molar mass of a compound is the sum of the

masses of the elements in its formula:

-Eg 1:  $C_2H_6O$  (C=12.0, H=1.0, O=16.0)  
Molar mass =  $12.0 \times 2 + 1.0 \times 6 + 16.0 \times 1 = 46.0g$   
The number of moles of a substance present in a given mass is given by:

Moles =  $\frac{\text{mass used}}{\text{molar mass}}$   
Eg: How many moles of ethanol ( $C_2H_6O$ ) are present in 69.0 g?

Moles =  $\frac{\text{mass used}}{\text{molar mass}} = \frac{69.0}{46.0} = 1.5 \text{ mol}$

#### ATOMIC STRUCTURE

Atoms are made of:

Protons: mass = 1, charge = +1

Neutrons: mass = 1, charge = 0

Electrons: mass = 0, charge = -1

Is a square on the periodic table the smaller number, the proton number gives the number of protons or electrons and the bigger number, the nucleon number the number of protons and neutrons together.



Eg 1: Boron has 5 protons, 6 neutrons, 5 electrons



Eg 2: Phosphorus has 15 protons, 16 neutrons 15 electrons

# Chemistry Revision Guide For Igcse Coordinated Science

**Joacim Rocklöv**



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## **Table of Contents Chemistry Revision Guide For Igcse Coordinated Science**

1. Understanding the eBook Chemistry Revision Guide For Igcse Coordinated Science
  - The Rise of Digital Reading Chemistry Revision Guide For Igcse Coordinated Science
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Revision Guide For Igcse Coordinated Science
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Revision Guide For Igcse Coordinated Science
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Revision Guide For Igcse Coordinated Science
  - Personalized Recommendations
  - Chemistry Revision Guide For Igcse Coordinated Science User Reviews and Ratings
  - Chemistry Revision Guide For Igcse Coordinated Science and Bestseller Lists

5. Accessing Chemistry Revision Guide For Igcse Coordinated Science Free and Paid eBooks
  - Chemistry Revision Guide For Igcse Coordinated Science Public Domain eBooks
  - Chemistry Revision Guide For Igcse Coordinated Science eBook Subscription Services
  - Chemistry Revision Guide For Igcse Coordinated Science Budget-Friendly Options
6. Navigating Chemistry Revision Guide For Igcse Coordinated Science eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Revision Guide For Igcse Coordinated Science Compatibility with Devices
  - Chemistry Revision Guide For Igcse Coordinated Science Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Revision Guide For Igcse Coordinated Science
  - Highlighting and Note-Taking Chemistry Revision Guide For Igcse Coordinated Science
  - Interactive Elements Chemistry Revision Guide For Igcse Coordinated Science
8. Staying Engaged with Chemistry Revision Guide For Igcse Coordinated Science
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Revision Guide For Igcse Coordinated Science
9. Balancing eBooks and Physical Books Chemistry Revision Guide For Igcse Coordinated Science
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Revision Guide For Igcse Coordinated Science
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Revision Guide For Igcse Coordinated Science
  - Setting Reading Goals Chemistry Revision Guide For Igcse Coordinated Science
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Revision Guide For Igcse Coordinated Science
  - Fact-Checking eBook Content of Chemistry Revision Guide For Igcse Coordinated Science
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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