

44

- known as: (A) cast iron (B) pig iron
(C) wrought iron (D) steel
(xvi) A coordinate covalent bond is always formed between the two: (A) like atoms (B) unlike atoms
(C) similar atoms (D) like and unlike atoms
(xvii) The first member of alkene series is:
(A) Methane (B) Ethane (C) Ethene (D) Ethyne

CHEMISTRY

2010

Time: 2 ½ Hours

Max. Marks: 68

SECTION "B" (SHORT-ANSWER QUESTIONS)(42)

NOTE: Answer 14 questions from this section.

2. Name three scientists of the modern period and write one contribution of each of in the field of Chemistry.
3. Define the following types of chemical reactions and give one example of each of them with a balanced chemical equation:
(i) Decomposition reaction
(ii) Single Displacement reaction
(iii) Combustion reaction
4. Balance the following equations:
(i) $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$ (ii) $\text{KNO}_3 \rightarrow \text{KNO}_2 + \text{O}_2$
(iii) $\text{NH}_3 + \text{O}_2 \rightarrow \text{NO}_2 + \text{H}_2\text{O}$
5. Name the three types of radioactive rays and write one characteristic of each.
6. Define the following:
(i) Fusion (ii) Sublimation (iii) Diffusion
7. Define Molarity and calculate the molarity of a solution that contains 20gm of NaOH in 500ml of the solution. (Atomic Masses: Na = 23, O = 16, H = 1)
8. Define the following:
(i) Ampere (ii) Coulomb (iii) Equivalent Mass
9. State Bronsted-Lowry concept of acids and bases and explain it with the example of HCl (as an acid) and H_2O (as a base).
10. Write three important uses of Epsom salt.
11. Define a Mole. Calculate no. of moles in 50gm of H_2O .
12. Write the chemical formulae of the following substances.
(i) Washing soda (ii) Baking soda (iii) Caustic soda
13. Define Exothermic and Endothermic reactions and write one example of each with a balanced chemical equation
14. Write chemical reaction of water with the following elements with balanced chemical equations:

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