

Chemical reactions involve changing the collection from one state of matter to another. However, during a chemical reaction, matter is not destroyed or created. The mass of the system (matter) is the same as the mass of the system before the reaction. This is the law of conservation of mass.

Recall that chemical reactions involve the breaking of bonds. These bonds are in the molecules and chemical reactions do not destroy the molecules, but they rearrange the atoms. Therefore, there are certain reactions that do involve changing the bonds. These chemical reactions involve the breaking of bonds and the formation of new bonds.

1. Fill in the table below as a reaction. Fill in the table with the particles that are involved in the reaction. The table is for a reaction between the elements. It gives the number of particles and the number of atoms. The table is for a reaction between the elements. It gives the number of particles and the number of atoms.

Reaction	Total Particles Before Reaction	Total Particles After Reaction	Total Atoms Before Reaction	Total Atoms After Reaction	Total Particles Before Reaction	Total Particles After Reaction
$H_2 + O_2$	2	2	2	2	2	2
$2H_2 + O_2$	3	3	6	6	3	3
$2H_2 + O_2 \rightarrow 2H_2O$	3	2	6	6	3	2
$2H_2 + O_2 \rightarrow 2H_2O$	3	2	6	6	3	2

2. Fill in the table below as a reaction. Fill in the table with the particles that are involved in the reaction.

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RD Boyd



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