

Binary Ionic Compounds

- Made up of _____ elements.
- Name the _____ ion (the _____).
- Name the _____ ion (the _____), a charged ion.
ending in _____.
- Always make the ion (anion) _____ than the cation and _____ ion (is a compound, not a _____ compound).
in parentheses after the name of the _____.
to show the _____.

Example: NaCl sodium chloride
 Cu_2S copper sulfide

AgCl magnesium chloride
 SnCl_2 tin chloride

Tertiary Ionic Compounds

- Made up of _____ elements.
- Name the _____ then name the _____ polyatomic anion.
_____ will have a charge that ending in "ide".

Example: $\text{Fe}_2(\text{SO}_4)_3$ cobalt sulfate
 FeCl_2O_4 iron(II) chromate

Chemistry Note Taking Guide Episode 603

Patrick Vollmar



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