

Ultimate Chemistry Unit One Notebook

Classification of Matter



Pure Substances

- Pure substances cannot be broken down into different substances.
- They are uniform and constant in composition

Elements

- Atomic Elements: only one atom in the formula
Example: Au, He, Na
- Molecular Elements: an element that exists as a molecule; there are two or more of the same atoms bonded together
Example: O₂, H₂, Br₂, Cl₂, I₂, F₂, N₂, O₃, P₄, S₈

Compounds

- Ionic Compound: a compound composed of a positively charged ion and a negatively charged ion
Example: NaCl, Sn(SO₃)₄
- Molecular Compound: a compound composed of two or more non-metal elements
Example: H₂O, CO₂

Mixtures

- Mixtures are made of two or more substances and can be separated by physical means
- Homogeneous mixtures are mixtures that are uniform throughout. A solution is a special type of homogeneous mixture.
Example: salt water
- Heterogeneous mixtures are mixtures that are not uniform throughout. Also known as mechanical mixtures
Example: soup

Chemistry Neutralization Guide

Judd E. Hollander



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