

**6.5 Multiple Choice Questions on Spectrometric Methods of Analysis**

1. The Beer-Lambert law states that  $A = \epsilon cl$ . The units of  $\epsilon$  are
  - (a)  $\text{cm mol L}^{-1}$
  - (b)  $\text{L mol}^{-1} \text{cm}^{-1}$
  - (c) none – the quantity is dimensionless
  - (d) none of the above
2. If the spectrum of a compound is recorded at different concentrations, which of the following values will change?
  - (a) Absorbance
  - (b)  $\lambda_{\text{max}}$
  - (c)  $\epsilon$
  - (d) Absorbance and the  $\lambda_{\text{max}}$
3. The wavenumber of a transition is  $2000 \text{ cm}^{-1}$ . To what part of the EM spectrum does this belong?
  - (a) IR
  - (b) Microwave
  - (c) Radio wave
  - (d) UV-vis
4. What is the absorbance of a  $0.1 \text{ mol dm}^{-3}$  aqueous solution of a nickel(II) salt which has  $\epsilon = 2.95 \text{ dm}^3 \text{ mol}^{-1} \text{ cm}^{-1}$ , and the path length of the solution cell used for the measurement is  $1 \text{ cm}$ ?
  - (a) 0.295
  - (b) 0.340
  - (c) 29.5
  - (d) 0.029
5. A shift to a lower wavenumber for an absorption peak in a spectrum corresponds to
  - (a) a shift to a higher energy
  - (b) a shift to a lower frequency
  - (c) a shift to a higher frequency
  - (d) a shift to a lower wavelength
6. Which of the following bonds will give the strongest absorption in the IR region?
  - (a) S-H
  - (b) N-H
  - (c) H-H
  - (d) O-H

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