

UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

- A. 1130 kJ/mol B. 565 kJ/mol **C. 283 J/mol** D. 300 kJ/mol E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine (N_2H_4) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.

- A. 95.0 kJ/mol** B. 2.96 kJ/mol C. 3.55×10^3
D. 92.4 J/mol E. 1.13×10^5 kJ/mol

3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone
- O_3
- .



Given the following data:



- A. -1083.7 kJ B. -693.8 kJ C. -198.8 kJ **D. -93.7 kJ** E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?

- A. $q < 0$, $\Delta E > 0$ **B. $q < 0$, $\Delta E < 0$** C. $q > 0$, $\Delta E > 0$
D. $q > 0$, $\Delta E < 0$ E. $q > 0$, $\Delta E = 0$

5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?

- A. 19.1 kJ B. 874 kJ **C. 1.91×10^{-2} kJ**
D. 3.35×10^5 kJ E. 335 kJ

6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.



- A. -90.7 kJ **B. -98.3 kJ** C. -49.2 kJ D. -108.4 kJ E. -196.6 kJ
7. When Karl Kaverman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.

- A. 2.9 °C B. 14 °C **C. 22 °C** D. 25 °C E. 47 °C

Chemistry Multiple Choice Questions And Answers

J Elliott



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needs water acidity or alkalinity acids properties and reactions amphoteric oxides basic acidic neutral and amphoteric chemical formulas chemical reactions chemistry reactions college chemistry mineral acids general properties neutralization ordinary level chemistry organic acid pH scale acid and alkali properties bases and reactions strong and weak acids and universal indicators The chapter Chemical Bonding and Structure MCQs covers topics of ions and ionic bonds molecules and covalent bonds evaporation ionic and covalent substances ionic compounds crystal lattices molecules and macromolecules organic solvents polarization and transfer of electrons The chapter Chemical Formulae and Equations MCQs covers topics of chemical formulas chemical equations atomic mass ionic equations chemical reactions chemical symbols mixtures and compounds molar mass percent composition of elements reactants relative molecular mass valency and chemical formula and valency table The chapter Electricity MCQs covers topics of chemical to electrical energy applications of electrolysis reactions conductors and non conductors dry cells electrical devices circuit symbols electrolytes non electrolytes organic solvents polarization and valence electrons The chapter Electricity and Chemicals MCQs covers topics of chemical to electrical energy dry cells electrolyte non electrolyte and polarization The chapter Elements Compounds and Mixtures MCQs covers topics of elements compounds mixtures molecules atoms and symbols for elements

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Table of Contents Chemistry Multiple Choice Questions And Answers

1. Understanding the eBook Chemistry Multiple Choice Questions And Answers
 - The Rise of Digital Reading Chemistry Multiple Choice Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Multiple Choice Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Multiple Choice Questions And Answers
 - User-Friendly Interface

4. Exploring eBook Recommendations from Chemistry Multiple Choice Questions And Answers
 - Personalized Recommendations
 - Chemistry Multiple Choice Questions And Answers User Reviews and Ratings
 - Chemistry Multiple Choice Questions And Answers and Bestseller Lists
5. Accessing Chemistry Multiple Choice Questions And Answers Free and Paid eBooks
 - Chemistry Multiple Choice Questions And Answers Public Domain eBooks
 - Chemistry Multiple Choice Questions And Answers eBook Subscription Services
 - Chemistry Multiple Choice Questions And Answers Budget-Friendly Options
6. Navigating Chemistry Multiple Choice Questions And Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Multiple Choice Questions And Answers Compatibility with Devices
 - Chemistry Multiple Choice Questions And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Multiple Choice Questions And Answers
 - Highlighting and Note-Taking Chemistry Multiple Choice Questions And Answers
 - Interactive Elements Chemistry Multiple Choice Questions And Answers
8. Staying Engaged with Chemistry Multiple Choice Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Multiple Choice Questions And Answers
9. Balancing eBooks and Physical Books Chemistry Multiple Choice Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Multiple Choice Questions And Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Multiple Choice Questions And Answers
 - Setting Reading Goals Chemistry Multiple Choice Questions And Answers
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Chemistry Multiple Choice Questions And Answers
 - Fact-Checking eBook Content of Chemistry Multiple Choice Questions And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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