

CHAPTER 4

MiniLab 4 – Model Isotopes

Analysis

1. The relative number of pre- and post-1982 pennies in the bag determines the percentage abundance of each group.
2. The atomic mass of a penny depends upon the mixture of pennies each student receives. Sample data is shown here.
 $\text{mass contribution (pre-1982)} = (55.0\%)(3.11 \text{ g}) = 1.71 \text{ g}$
 $\text{mass contribution (post-1982)} = (45.0\%)(2.55 \text{ g}) = 1.15 \text{ g}$
 $\text{atomic mass} = (1.71 \text{ g} + 1.15 \text{ g}) = 2.86 \text{ g}$
3. A different mixture would have a different relative abundance and a different atomic mass.
4. Masses of individual pennies will vary due to wear.

Expected Results:

Pre-1982 pennies have a greater mass than post-1982 pennies.

Mass of ten pre-1982 pennies = 31.10 g

Average mass of a pre-1982 penny = 3.11 g

Mass of ten post-1982 pennies = 25.48 g

Average mass of a post-1982 penny = 2.55 g

The atomic mass depends on the mixture analyzed.

ChemLab 4 – Calculate the Atomic Mass of the Element “Snackium”

Pre-Lab

2. protons, neutrons, and electrons
3. Answers will vary.
4. If there are more heavy subatomic particles, the isotope will be heavier. If there are fewer heavy subatomic particles, the isotope will be lighter.

Analyze and Conclude

1. Answers will vary depending on the type of snack bags.
2. Answers will vary depending on the type of snack bags. If the snacks all have similar masses, the average atomic mass will be closest to the most abundant snack.

3. Answers will vary. The students should find that the atomic mass is closest to the most abundant piece. In the sample data, the atomic mass is closer to the chex pieces than the pretzels or bagel chips due to their high percent.
4. Answers will vary. Students should realize that the small sample size and difference in samples will lead to differences.
5. The atomic mass is different from the mass number because it is an average and therefore will not be a whole number.
6. Answers will vary depending on the reference sources used. They should find most elements have numerous natural isotopes.
7. The error in the lab is due to small sample size. The experiment could be modified by starting with a much larger sample size.

Inquiry Extension

Answers will vary. The students should be able to predict the most abundant element based on the atomic mass listed on the periodic table.

Teaching Transparency 11 – Cathode Ray Experiments

1. A cathode ray is a beam of negatively charged particles or electrons.
2. Answers may vary. All three experiments have a voltage source, a cathode, an anode, gas at low pressure, and a phosphor that allows the position of the cathode ray to be determined.
3. The cathode ray travels from the cathode toward the anode, passing through the hole in the anode and traveling in a straight line through the tube, where it strikes the phosphor screen.
4. In B, the cathode ray passes through a magnetic field, whereas in C, the cathode ray passes between two electrically charged plates. Both experiments show that the cathode ray is made of charged particles.
5. The experiment shows that the cathode ray consists of charged particles that are affected by a magnetic field.
6. The negatively charged cathode ray bends toward the positively charged plate due to electrical forces of attraction.

Chemistry Matter And Change Resource Answers

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