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PRECIPITATION REACTION LAB

Introduction:

The majority of ionic solids like salts (KCl and NaCl) are soluble in water because the polar water molecules surround the individual ions of the salt. Those that do not dissolve and go into solution form solid products called precipitates. These precipitates have many colors and often help scientists identify what the precipitate is present. In meteorology the term precipitation refers to meteorological phenomena such as rain or snow. Precipitation is chemical solutions occurs when two chemicals react to form a product that is insoluble in water and falls out of solution like rain or snow. A precipitate is a solid substance that separates from solution during a chemical reaction. A precipitate can be identified by the cloudy, milky, gelatinous, or grainy appearance it gives to the mixture.

A barium sulfate precipitate can be produced by the reaction of barium chloride and sodium sulfate. A chemical equation to describe the reaction is written and balanced like this:

$$BaCl_{2 (aq)} + Na_2SO_{4 (aq)} \rightarrow 2 NaCl_{(aq)} + BaSO_4(s)$$

Barium sulfate, BaSO₄ (s), is a common precipitate used as an X-ray contrast medium because it is insoluble in water an opaque to x-rays. Typically a patient drinks an aqueous slurry of barium sulfate just before he is x-rayed. The precipitate coats his stomach and intestines. These organs then show up on the x-ray film in vivid contrast, aiding the doctor's diagnosis. Notice the reaction that forms BaSO₄ is a double-replacement reaction in which the cations and anions of

the reactants trade partners to form the products.

To write a total ionic equation, rewrite all aqueous substances as their component ions and keep all solid substances unchanged.

$$Ba^{+2}_{(aq)} + 2Cl^{-}_{(aq)} + 2Na^{+}_{(aq)} + SO_4^{-2}_{(aq)} \rightarrow 2Na^{+}_{(aq)} + 2Cl^{-}_{(aq)} + BaSO_{4(s)}$$

Cancel all components of the reaction that are identical on both the reactant and product side of the reaction. These are called spectator ions. Write the net ionic equation by rewriting the reaction equation including only those reactants and products that change in the reaction.

Chemistry Lab Report Precipitation Reactions Answers

A. K. Raj

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Laboratory Manual for Science \(\begin{aligned} \begin{alig 10 These are complimentary to the Science textbooks of the respective classes The manuals cover a wide range of age appropriate experiments that give hands on experience to the students. The experiments help students verify scientific truths and principles and at the same time expose them to the basic tools and techniques used in scientific investigations Our manuals aim not only to help students better comprehend the scientific concepts taught in their textbooks but also to ignite a scientific quest in their young inquisitive minds Core Science Lab Manual with Practical Skills for Class IX V. K. Sally, Chhaya Srivastava, Goyal Brothers Prakashan, 2019-01-01 Goyal Brothers Prakashan Chemistry: Inorganic Qualitative Analysis in the Laboratory Clyde Metz, 2012-12-02 Chemistry Inorganic Qualitative Analysis in the Laboratory is a textbook dealing with qualitative analysis in the laboratory as well as with the process of anion and cation analysis. The book presents an overview of the subject of inorganic qualitative analysis including as the equipment reagents and procedures that are going to be used in the laboratory Preliminary experiments include the classification of precipitates handling precipitates separation techniques flame tests Brown ring test solvent extraction The text also describes in detail how to prepare the experiment for anion and cation analysis such as testing for water solubility in a solid sample or the sodium carbonate treatment of a water soluble sample The book also explains the qualitative analysis for anions in preliminary and specific tests In the qualitative analysis for cations the student follows different procedures for Cation Groups I II III IV or V For example the ions of Cation Group V cannot be precipitated by any Cation Groups I IV reagents nor by any single group reagent The textbook is suitable for both chemistry teachers and freshmen students **Chemistry in the Laboratory James** M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations Practical Skills in Science Class 09 R.P. Manchanda, Practical Book The Chemical Gazette, Or, Journal of Practical Chemistry, in All Its Applications to Pharmacy, Arts, and Manufactures William Francis, Henry Croft, 1858 The Chemical Gazette, Or, Journal of Practical Chemistry, in All Its Applications to Pharmacy, Arts, and Manufactures ,1858 Applications of Environmental Aquatic Chemistry Eugene R. Weiner, 2012-12-07 Professionals and students who come from disciplines other than chemistry need a concise yet reliable guide that explains key concepts in environmental chemistry from the fundamental science to the necessary calculations for applying them Updated and reorganized Applications of Environmental Aquatic Chemistry A Practical Guide Third Edition provides the essential background for understanding and solving the most frequent environmental chemistry problems Diverse and self contained chapters offer a centralized and easily navigable framework for finding useful data

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