

Kinetic Properties of a Chemical Reaction

INTRODUCTION

Beginning chemistry students sometimes have the mistaken idea that all chemical reactions happen instantly. This is probably because they have seen reactions such as fire or the precipitation of silver chloride, which appear to be almost instantaneous. However, one need only think of the rusting of an iron nail or the fermenting of fruit juice in wine making, which are very slow reactions, to realize that this impression is incorrect. As the speed or rate of reactions differ so greatly, therefore, the study of chemical rates or kinetics is a very important discipline. In this experiment, you will examine various aspects of the kinetics of a so-called "clock reaction":



colorless colorless colorless brown

In this experiment, you will investigate the effects upon the rate of three different factors:

- concentration of reactants;
- temperature; and
- the presence of a catalyst.

In order to accomplish this, you will perform the reaction four times, under different conditions, monitoring the progress of the reaction by observing the rate at which the brown triiodide ion appears. The four different sets of conditions are:

- at 30 °C;
- at 30° C with reduced amount of $\text{S}_2\text{O}_8^{2-}$;
- at 37 °C; and
- at 37 °C with a catalyst present.

Thus, you will be able to observe the effect of changing concentration by comparing the results from trials (i) and (ii); the effect of temperature by comparing trials (i) and (iii); and the effect of a catalyst by comparing runs (iii) and (iv).

As this is a chemistry class, it is not sufficient just to observe the various effects qualitatively; we will want to express the results numerically. We shall accomplish this by measuring the concentration of the triiodide anion as time progresses, and converting these data to useable numbers. In order to measure the iodide ion concentration, we shall use the fact that the ion is colored—thus, the higher the concentration of iodide in a particular solution, the darker the color. Things appear colored because they absorb particular wavelengths of visible light and allow other wavelengths to pass through (to be transmitted). More intense-

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