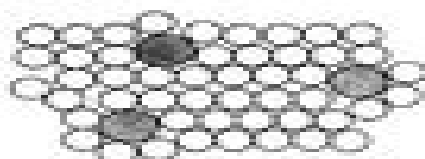


15. The diagram below shows the structure of an alloy.



Which statement about alloys is correct?

- A The alloy brass has a chemical formula.
 - B High carbon steel alloys are soft and easily shaped.
 - C Alloys can only be formed by mixing copper or iron with other metals.
 - D In an alloy there is attraction between positive ions and a 'sea of electrons'.
16. Calcium carbonate, CaCO_3 , reacts with an acid, HA as shown below.
- $$\text{CaCO}_3(\text{s}) + 2\text{HA}(\text{aq}) \rightarrow \text{CaA}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$$
- What is the minimum mass of acid required to react completely with 50 g of calcium carbonate? [M: HA, 46; CaCO_3 , 100]
- A 4.6 g
 - B 9.2 g
 - C 10 g
 - D 20 g
17. When concentrated sulfuric acid reacts with sodium iodide the products include sulfur, iodine, hydrogen sulfide and sulfur dioxide. Which statement is correct?
- A Hydrogen sulfide is the product of a reduction reaction.
 - B Iodide ions are stronger oxidising agents than sulfate ions.
 - C Sulfur atoms from the sulfuric acid are both oxidised and reduced.
 - D Sulfur atoms from the sulfuric acid are oxidised to make sulfur dioxide.
18. Attaching pieces of magnesium to underground iron pipes can protect the iron from corrosion. Which reaction protects the iron from corrosion?
- A $\text{Fe}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Fe}(\text{s})$
 - B $\text{Fe}(\text{s}) \rightarrow \text{Fe}^{2+}(\text{aq}) + 2\text{e}^-$
 - C $\text{Mg}^{2+}(\text{aq}) + 2\text{e}^- \rightarrow \text{Mg}(\text{s})$
 - D $\text{Mg}(\text{s}) \rightarrow \text{Mg}^{2+}(\text{aq}) + 2\text{e}^-$

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Mark Steyvers



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