

Solubility Rules:-NO₃⁻¹ all soluble-Cl⁻ all soluble except AgCl, Hg₂Cl₂ and PbCl₂-SO₄⁻² all soluble except BaSO₄, PbSO₄ and SrSO₄-OH⁻ all insoluble except Group 1, Ba(OH)₂, Sr(OH)₂, Ca(OH)₂-CO₃⁻² all insoluble except NH₄⁺ and Group 1-S⁻² all insoluble except Group 1 and Group 2 and NH₄⁺**Polyatomics:**-Acetate (C₂H₃O₂⁻²)-Chlorate (ClO₃⁻¹)-Chlorite (ClO₂⁻¹)-Cyanide (CN⁻¹)-Hydroxide (OH⁻¹)-Nitrate (NO₃⁻¹)-Nitrite (NO₂⁻¹)-Nitrite (NO₂⁻¹)-Carbonate (CO₃⁻²)-Chromate (CrO₄⁻²)-Dichromate (Cr₂O₇⁻²)-Oxalate (C₂O₄⁻²)-Sulfate (SO₄⁻²)-Sulfite (SO₃⁻²)-Arsenate (AsO₄⁻³)-Arsenite (AsO₃⁻³)-Phosphate (PO₄⁻³)-Phosphite (PO₃⁻³)-Ammonium (NH₄⁺¹)-Bicarbonate (HCO₃⁻¹)**Acids:**

-HCl = hydrochloric

-H₂SO₄ = sulfuric-HNO₃ = nitric-CH₃COOH = acetic-H₃PO₄ = phosphoric-H₂CO₃ = carbonicMatter is the physical material of the universe and exists in 3 states: solid, liquid, gasSubstance is matter that has a uniform and definite compositionMixture is a physical blend of substancesHomogenous Mixture has uniform composition throughout (solution) Heterogeneous Mixture does notPhysical Properties: observed or measuredChemical Properties: ability to undergo chemical changeLaw of Conservation of Mass: matter/mass can be neither created nor destroyedExtensive Properties: mass and volumeIntensive Properties: compositionSolid: definite shape, volume, not easily compressed Liquid: indefinite shape, definite volume, not easily compressedGas: indefinite shape, volume, easily compressed *Particles are always moving in S, L, GIndications of Chemical Reaction: dissolving, releasing/absorption of heat, production of gas, color change, precipitateQualitative give descriptions, no numbersQuantitative number-basedAccuracy is how close to a single measurementPrecision is how close several measurements are together

Error = |expt - accepted| % Error = |expt-accepted|/accepted * 100 Percent Yield = expt/theoretical * 100

1 kilo(unit) = 1000 (unit) 1 (unit) = 1000 milli

1 (unit) = 100 centi

micro=10⁶nano=10⁹

K = C + 273

C = K - 273

H₂O FP=273K

BP=373K Density = mass/volume

Density of H₂O = 1 g/mLSig Figs: For addition & subtraction use least # of decimal places

/ For Multiplication & Division use least # of sig figs

Dalton's Atomic Theory: 1) all elements are composed of atoms

2) atoms of same element are identical

3) atoms can physically mix/chemically combine in whole #d ratios

4) chemical rxns happen; atoms don't change

Proton = pos., change = radioactivityNeutron = neutral, change = isotopeElectron = neg., change = ion

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Nuclear Science Abstracts ,1975 [Handbook of dietary and nutritional aspects of bottle feeding](#) Victor R. Preedy, Ronald Ross Watson, Sherma Zibadi, 2023-08-07 Newborn babies are usually fed via the breast which ensures optimal development and emotional mother child bonding However in some circumstances breast feeding may be either inadequate or impractical Historically deficiencies in the provision of breast milk were addressed by wet nurse feeding but this was superseded by the introduction of artificial milk feeds i e formula or bottle feeds This handbook covers formula feeds in the most comprehensive way Each contribution starts with convenient summary points Six parts give information about historical and international aspects South American African and Chinese perspectives composition microbiological and chemical contaminants allergy and immunology effects on physical development and metabolic responses and effects on psychological and neurodevelopment The Handbook of dietary and nutritional aspects of bottle feeding is essential reading material for nutritionists dieticians paediatricians midwives pharmacologists health care professionals general practitioners and those interested in babies health in general *Report of the Surgeon-General of the Army to the Secretary of War for the Fiscal Year Ending ...* United States. Surgeon-General's Office, 1936 Provides data statistical and tabular on the operations and

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