CHEMISTRY: COUNTING ATOMS IN COMPOUNDS WORKSHEET #7.0.1

INSTRUCTIONS: Write the quantity of atoms of each element opposite the formula of the compound for the quantity of formula units and molecules shown:

For example: $5P_2O_3$ $P = (5 \times 2 =) 10$ $O = (5 \times 3 =) 15$

For example: $4Zn(NO_3)_2$ $Zn = (4 \times 1 =)4$ $N = (4 \times 1 \times 2 =)8$ $O = (4 \times 3 \times 2 =)24$

- 1. 4K₂CO₃ K = ____ C = ___ O = ___
- 2Sr₃(PO₄)₂ Sr = ____ P = ____ O = ____
- 3N₄O₁₀ N = _____ O = ____
- 2(NH₄)₃N N = ____ H = ____
- 8Cl₂O Cl = ___ O = ___
- 6. Ca(C₂H₃O₂)₂ Ca = ___ C = ___ H = __ O = ___
- 7. 12NaBr Na = ____ Br = ____
- 8. 4Al(OH)₃ Al = ___ O = ___ H = ___
- 9. 3NaHCO₃ Na = ___ H = ___ C = ___ O = ___
- 5Ga₂(Cr₂O₇)₃ Ga = ____ Cr = ___ O = ____
- 11. 7C₂S₂ C = ____ S = ____
- 12. 4Fe₂O₃ Fe = ____ O = ____
- 13. 6Ba(MnO₄)₂ Ba = ____ Mn = ___ O = ____
- 14. 3V₂O₅ V = ____ O = ____
- 15. 2KNO₃ K = ____ N = ___ O = ____
- 9MgSO₄ Mg = ____ S = ____ O = ____
- 17. 5Al₂(SiO₃)₂ Al = ____ Si = ___ O = ____
- 18. 4Au(IO₃)₃ Au = ____ I = ____ O = ____

(Continued)

Chemistry Counting Atoms In Compounds

Darrell D. Ebbing, Albert E. Richardson

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