

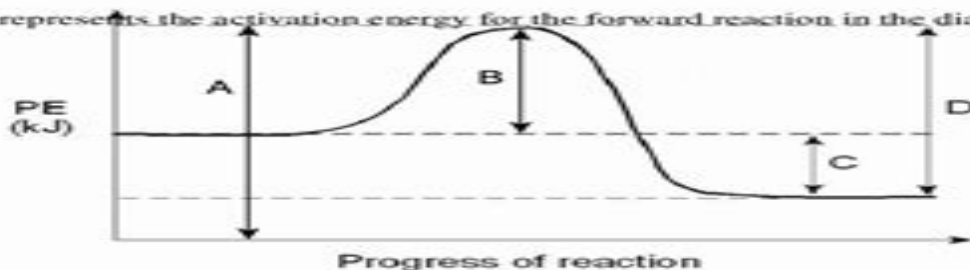
PART I
Total Value: 50%

Instructions: Shade the letter of the correct answer on the computer scorable answer sheet provided.

1. Why does increasing temperature increase reaction rate?

- (A) decreases activation energy
(B) decreases the percentage of effective collisions
✓ (C) increases activation energy
(D) increases the percentage of effective collisions

2. Which represents the activation energy for the forward reaction in the diagram below?



- ✓ (A) A
(B) B
(C) C
(D) D

3. Which type of change is best for monitoring the reaction rate of the reaction below?



- (A) colour
✓ (B) mass
(C) pH
(D) volume

4. Which describes an activated complex?

- (A) has a lower potential energy than products
(B) has a lower potential energy than reactants
✓ (C) stable species that forms upon collision of reactants
(D) unstable species that forms upon collision of reactants

5. According to the collision theory, why does the reaction below occur in more than one step?



- ✓ (A) low $\text{C}_5\text{H}_{12}(\text{g})$ concentration
(B) low probability of a nine-particle collision
(C) particles collide with insufficient kinetic energy
(D) temperature is very high

Chemistry Collision Theory Study Guide Answers

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**



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