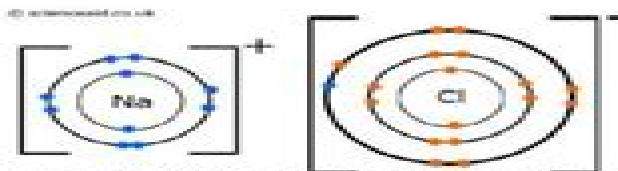


Chemical Bonding

Ionic Bonding

Ionic Compounds: Composed of anions and cations, these compounds do not form molecules. The chemical formulas for these are in the simplest ratio. Components are combined so that amount of positive and negative charge is equal. More likely than not forms a crystalline solid.



Think of it as an awkward first date between two middle schoolers: *no chemistry (so punny) no compound!*

*Ionic compounds form so that each atom, by **gaining or losing** electrons, has an octet of electrons in its highest occupied energy level.*

Formation: Metals lose electrons to form positively-charged cations, whereas nonmetals gain electrons to form negatively-charged anions.

The goal of the formation of a compound can be seen through electron configurations. Whatever adds up to make that of a noble gas (electron configuration) is desired to stabilize.

Ions come together due to electrostatic attraction, since opposites attract.

Formation (continued):

Attractive Forces: (What is attracted together) Oppositely charged ions, nuclei and electron clouds of adjacent ions.

Repulsive Forces: (What is **not** attracted together) Like-charged ions, electrons of adjacent ions.

*The distance between the ions creates a balance between those forces; ions minimize their potential energy by combining in an arrangement called a **crystal lattice**. The specific lattice pattern created depends on the charges and*

Chemistry Chemical Bonding Study Guide Practice

Tyler L. Jackson



Chemistry Chemical Bonding Study Guide Practice:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Study Guide to Accompany Basics for Chemistry Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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