

Preparing for the Organic Chemistry Exam?

Atomic Structure

Atoms are made of three subatomic particles: Electrons, Protons, and Neutrons.

Electrons: e^- , mass 9.1×10^{-31} kg, charge -1.6×10^{-19} C.

Protons: p^+ , mass 1.67×10^{-27} kg, charge $+1.6 \times 10^{-19}$ C.

Neutrons: n^0 , mass 1.67×10^{-27} kg, charge 0.

Atomic Number (Z): Number of protons in the nucleus.

Mass Number (A): Sum of protons and neutrons in the nucleus.

Isotopes: Atoms of the same element with same atomic number but different mass numbers.

Atomic Structure (Continued)

Electron Configuration: The arrangement of electrons in different shells or orbitals.

Shells: K, L, M, N, O, P, Q.

Orbitals: Regions where the probability of finding an electron is maximum.

Quantum Numbers: Numbers that describe the energy, size, and shape of an orbital.

Pauli Exclusion Principle: No two electrons in an atom can have the same set of four quantum numbers.

Atomic Structure (Continued)

Atomic Structure of Matter: Matter is made of atoms and molecules.

Atoms: Smallest particles of an element that cannot be created or destroyed.

Molecules: Two or more atoms of the same or different elements joined together.

Compounds: Two or more different atoms joined together in a fixed ratio.

Atomic Structure (Continued)

Atomic Structure of Matter (Continued)

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