

In your textbook, read about molarity and preparing molar solutions.

Read the following problem and then answer the questions.

An 85.0-mL aqueous solution contains 7.54 g iron(II) chloride ( $\text{FeCl}_2$ ). Calculate the molarity of the solution.

5. What is the mass of the solute?

7.54 g

6. What is the volume of the solution?

85.0 mL solution

7. Write the equation that is used to calculate molarity.

$$M = \frac{\text{moles of solute}}{\text{L solution}}$$

8. In what unit must the amount of the solute be expressed to calculate molarity?

moles

9. In what unit must the volume of the solution be expressed to calculate molarity?

Liter

10. Write the expression needed to convert the volume of the solution given in the problem to the volume needed to calculate molarity.

$$85.0 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 0.0850 \text{ L}$$

### Lesson 13.2 continued

11. What quantity must be used to convert the mass of the solute given in the problem to the amount calculate molarity?

molar mass of  $\text{FeCl}_2$  126.75 g/mol

12. Write the expression used to calculate the amount of solute

$$7.54 \text{ g FeCl}_2 \times \frac{1 \text{ mol}}{126.75 \text{ g}} = 0.0595 \text{ mol}$$

13. Calculate the molarity of the solution. Show all your work.

$$\underbrace{7.54 \text{ g} \times \frac{1 \text{ mol}}{126.75}}_{\text{moles of Solute}} \times \underbrace{\frac{1}{0.0850 \text{ L}}}_{\text{L of Solution}} = \textcircled{0.700 \text{ M}}$$

3SF

# Chemistry Chapter 13 Study Guide

**T. W. Graham Solomons, Craig B.  
Fryhle, Scott A. Snyder**



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