

## CHEMICAL QUANTITIES

## **Practice Problems**

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

#### SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

- What is the molar mass of sucrose (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>)?
- 2. What is the molar mass of each of the following compounds?
  - a, phosphorus pentachloride (PCI-)
  - b. uranium hexafluoride (UF<sub>4</sub>)
- 3. Calculate the molar mass of each of the following ionic compounds:
  - a. KMinO<sub>4</sub>
  - b. Cas(PO<sub>4</sub>):
- 4. How many moles is 3.52 × 1024 molecules of water?
- 5. How many atoms of zinc are in 0.60 mol of zinc?
- 6. What is the mass of 1.00 mol of exygen (O2)?

## SECTION 10.2 MOLE-MASS AND MOLE-VOLUME RELATIONSHIPS

- 1. What is the molar mass of each of the following compounds?
  - a. CaHagOa

b. NaHCOs.

e. Calling

d. KNH4SO<sub>4</sub>

- 2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide (PbO).

- d. 1.50 × 10<sup>-2</sup> mol molecular oxygen (O<sub>5</sub>)
- b. 0.75 mol hydrogen sulfide (H<sub>2</sub>S)

- e. 2.30 mol ethylene glycol (C<sub>2</sub>H<sub>2</sub>O<sub>2</sub>)
- e. 0.00100 mol silicon tetrahydride (SiH<sub>4</sub>).
- How many grams are in 1.73 mol of dinitrogen pentoxide (N<sub>2</sub>O<sub>2</sub>)?
- How many grams are in 0.658 mol of calcium phosphate [Ca<sub>2</sub>(PO<sub>4</sub>)<sub>2</sub>]?
- 5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide (NaBr).

- d, 0.00100 g monochloromethane (CH<sub>2</sub>CI)
- b. 13.5 g magnesium nitrate [Mg(NO<sub>1</sub>)<sub>2</sub>]
- e.  $1.50 \times 10^{-3}$  g propylene glycol [C<sub>3</sub>H<sub>4</sub>(OH)<sub>2</sub>]
- e. 1.02 g magnesium chloride (MgCl<sub>2</sub>)
- 6. A chemist plans to use 435.0 grams of ammonium nitrate (NH<sub>4</sub>NO<sub>5</sub>) in a reaction. How many moles of the compound is this?
- 7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine (C<sub>20</sub>H<sub>24</sub>N<sub>2</sub>O<sub>2</sub>). What mass, in grams, should the laboratory technician obtain in order to make the solution?
- 8. What is the volume at STP of 2.66 mol of methane (CHa) gas?
- 9. How many moles is 135 L of ammonia (NH<sub>3</sub>) gas at STP?

# **Chemistry Chapter 10 Study Guide Answer Key**

Barbara M. Castleberry, Mary E. Lunz, Bradford L. Sahl

## **Chemistry Chapter 10 Study Guide Answer Key:**

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## **Table of Contents Chemistry Chapter 10 Study Guide Answer Key**

- 1. Understanding the eBook Chemistry Chapter 10 Study Guide Answer Key
  - The Rise of Digital Reading Chemistry Chapter 10 Study Guide Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemistry Chapter 10 Study Guide Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Chapter 10 Study Guide Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemistry Chapter 10 Study Guide Answer Key
  - Personalized Recommendations
  - Chemistry Chapter 10 Study Guide Answer Key User Reviews and Ratings
  - Chemistry Chapter 10 Study Guide Answer Key and Bestseller Lists
- 5. Accessing Chemistry Chapter 10 Study Guide Answer Key Free and Paid eBooks
  - Chemistry Chapter 10 Study Guide Answer Key Public Domain eBooks
  - Chemistry Chapter 10 Study Guide Answer Key eBook Subscription Services

- o Chemistry Chapter 10 Study Guide Answer Key Budget-Friendly Options
- 6. Navigating Chemistry Chapter 10 Study Guide Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chemistry Chapter 10 Study Guide Answer Key Compatibility with Devices
  - Chemistry Chapter 10 Study Guide Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Chapter 10 Study Guide Answer Key
  - Highlighting and Note-Taking Chemistry Chapter 10 Study Guide Answer Key
  - Interactive Elements Chemistry Chapter 10 Study Guide Answer Key
- 8. Staying Engaged with Chemistry Chapter 10 Study Guide Answer Key
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Chapter 10 Study Guide Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Chapter 10 Study Guide Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Chapter 10 Study Guide Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Chapter 10 Study Guide Answer Key
  - Setting Reading Goals Chemistry Chapter 10 Study Guide Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Chapter 10 Study Guide Answer Key
  - Fact-Checking eBook Content of Chemistry Chapter 10 Study Guide Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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