

Some solids, while crystallizing out from their solutions, combine with a definite quantity of water, known as water of crystallisation or water of hydration.

This water of crystallisation is in loose chemical combination and can be driven off as steam by heating the crystals above 100°C , the vapours produced are condensed and the liquid formed is tested. This liquid gives test of water. In this process, the crystals lose their water and become **anhydrous**, i.e. all the water combined in the crystalline material is removed.

Experiment : To show that hydrated copper sulphate crystals (blue vitriol) contain water of crystallisation.

Procedure : Take some powdered copper sulphate crystals in a clean dry test tube, and clamp it in a tilted position (Fig. 3.11) so that the condensed water does not trickle or flow back. Gently heat the crystals.

Observation : Drops of colourless liquid condense on the cooler parts of the test tube, leaving behind a residue that is anhydrous (without water) and amorphous (non-crystalline), i.e. with no definite shape or structure.



Fig. 3.11 To show loss of water of crystallisation on heating

When blue crystals of hydrated copper sulphate are heated in a test tube, they turn into a white powder, which turns back into a blue solid when a few drops of water are added.



When preparing crystals of a hydrated salt, we must be careful not to heat them too strongly when drying them, otherwise the product is a dehydrated powder, not crystals.

Determination of water of crystallisation

Take a known weight of crystals in a china dish

and heat it above 100°C . When the weight of the residue becomes constant, stop heating.

Wt. of crystals at room temperature = a g

Wt. of crystals after heating at 101°C = b g

Wt. of water = $(a - b)$ g

Now a g of crystals contain water = $(a - b)$ g

$$\therefore \% \text{ of water of crystallisation} = \frac{a-b}{a} \times 100$$

3.9.2 Anhydrous substances

A substance is anhydrous if it contains no water, for example, salts lacking their water of crystallisation. The term 'Anhydrous' is most often applied to crystalline substances when the water of crystallisation is removed.

Anhydrous can also refer to the gaseous form of some concentrated solutions such as ammonia to distinguish it from its aqueous solution form.

Anhydrous solvents are used to perform certain chemical reactions that cannot proceed in presence of water. Examples of anhydrous substances.

- Table salt is anhydrous sodium chloride.
- Gaseous HCl is anhydrous.

Anhydrous salts can be obtained from hydrated salts by carefully removing the water of crystallisation from them. It can be done by

- Direct heating.
- Heating in dry and hot air.
- Heating under vacuum.
- By using dehydrating/desiccating agents such as warm concentrated sulphuric acid.

The crystalline shape of a substance is not necessarily the result of the presence of water of crystallisation. In fact, there are a number of crystalline solids that crystallize from water without holding any water of crystallisation.

Examples of crystalline substance which do not contain water of crystallisation :

Common salt (NaCl).

Nitre (KNO_3).

Sugar ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$).

Potassium permanganate (KMnO_4).

Ammonium chloride (NH_4Cl).

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