

Molecular Structure – Part 1

Purpose of the Experiment

To enhance understanding of the relationship between formula and structure of various chemical species.

Equipment

Orbit Molecular Building System Centers and Tubes:

Atom Centers*

Single – white

Linear (2-points) - black

Trigonal planar (3-points) - black

Tetrahedral (4-points) - black, green (dark and light), yellow, purple, blue, red

Trigonal bipyramidal (5-points) – black, blue, red, yellow, purple, green

Octahedral (6-points) – black, blue, yellow, green

*Color code for atom centers:

Black	Carbon	Yellow	Sulfur
Red	Oxygen	Purple	Phosphorus
Blue	Nitrogen	Green	Halogen
White	Hydrogen		

Bonds

Tubes – rigid green, flex white

INTRODUCTION

The atoms in molecules and many polyatomic ions are bound together by covalent bonds and are locked into a relatively fixed spatial arrangement with respect to one another. Covalent bonds consist of shared pairs of valence electrons. It is sometimes helpful to think of these shared pairs of valence electrons as the overlap of atomic orbitals. Bonds may be single, double, or triple (representing the sharing of one, two, or three pairs of electrons respectively). Many atoms also contain one or more pairs of valence shell electrons that are not involved in bonding, but nevertheless influence the overall shape of the molecule. These are called nonbonding pairs. It is very important for the understanding of chemical principles that students of chemistry be able to visualize the three-dimensional shapes of molecules and polyatomic ions. The shape of a

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