

POST LAB QUESTIONS

Use the space below for your calculations for Part A.

$$K_b: \text{NH}_4^+ = 1.8 \times 10^{-5}$$

$$K_a: \text{Acetic Acid} = 1.8 \times 10^{-5}$$

$$K_{a2}: \text{H}_2\text{SO}_4 = 1.2 \times 10^{-2}$$

$$K_b: \text{Sodium Acetate} = 5.575 \times 10^{-10}$$



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Chemistry 1212 Laboratory Manual Experiment 5

Jicheng Xie



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