

Question 1

0 out of 1 points

A first order reaction $A \rightarrow B + C$ has a rate constant of 0.0125 s^{-1} . If the initial concentration of A is 0.500 M , what is the concentration of A after 5 minutes? (Note: do not include units in your answer, do not use scientific notation, present your answer to 2 significant figures.)

Selected Answer: ☒ 25

Correct Answer: ☒ 0.0118

Answer range: 0.002 (0.0008 - 47- 0.0138)

Question 2

0 out of 1 points

The reaction of nitrogen dioxide with fluorine generates nitryl fluoride:



To determine the rate law for this reaction, a chemist performed several initial-rate experiments using different initial concentration, with the details tabulated below.

Experiment (#)	$[\text{NO}_2] \text{ (M)}$	$[\text{F}_2] \text{ (M)}$	$[\text{NO}_2\text{F}] \text{ (M)}$	Initial Rate (M/s)
1	1.0×10^{-2}	5.0×10^{-3}	0.10×10^{-2}	2.0×10^{-4}
2	2.0×10^{-2}	5.0×10^{-3}	0.10×10^{-2}	4.0×10^{-4}
3	2.0×10^{-2}	5.0×10^{-3}	1.00×10^{-2}	4.0×10^{-4}
4	2.0×10^{-2}	7.5×10^{-3}	0.10×10^{-2}	6.0×10^{-4}

Determine the rate constant for this reaction. (Note: do not include units in your answer, do not use scientific notation, present your answer to 2 significant figures.)

Selected Answer: ☒ 120

Correct Answer: ☒ 40

Answer range: 2 (35.0 - 42.0) -47-

Question 3

1 out of 1 points

Consider the following reaction with an equilibrium constant $K_c = 20.06$. Fill in the blanks with either less or greater.



From this information it can be conclude that the rate of the forward reaction is **[2] greater** than the rate of the reverse reaction and that the concentration of the reactants will be **[2] greater** than the concentration of the products.

Specified Answer: ☒ Great

Chemistry 1020 Lab Manual Answers

Yijin Wang



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