

CHEMISTRY Terminology The ABCs of chemistry

A: Symbol for atomic mass marrier

absolute zero: Lowest temperature; 0 Kelvin

Act Symbol for the element actinium, Z = 89.

accuracy: The closeness of measurement to true value.

acetate anion: CallaDy, charge -1

postic acid: C-01.O-

acetome: Directled ketome: organic solvent.

acid-base indicator: A chemical that changes color lift the pHI champes

acid-base titration: A method for determining acid or base concentrations.

noid definition: See Archenius acid, Brensted-

Lower sold, Lewis acid.

acid ionization constant (K.): Equilibrium constant for acid dissociation.

acidic solution; pit Buckeyer 1

actinide: Element switch 2: ~ 90 or allower: radiosective.

activation energy

(IE.)): A prescess is energy.

Resident

Е.

Activation Energy

Product

adhesion: Attraction of unlike molecules or

Age: Symbol for the element silver, Z = 47

Alt: Symbol for the element aluminum, Z = 13.

alcohol: Organic compound with -OH group; BOH aldehyde: Carboxyl group bonded to hydrogen and an organic group

alkali metal family: Lithium, sodium, potosium, nubidium, opsium, and francium; column #1

alkaline earth metal family: Beryllium, magnesium, calcium, strontium, barium, and radium; column #2.

allkame: Hydrocurbon; all C-C single bonds

alkene: Hydrocarbon; I or more C+C double bonds. alloyse: Hydrocarbon, I ar more CFC triple bonds.

allotropes: Two or more forms of an element

alticy: Solution of 2 or more metals.

alpha (a): Greek letter that denotes radioactive particle and various scientific variables.

alpha particle: Helium micleus; charge +2

Amc Sverbol for the element sovericium, Z = 95

amalgame Alloy of mercury with other inetal

amine: Organic base; RRRN; modified amountain

amino acid: Compound with organic acid and organic

base property; forms proteins and peptides

ammonia: Nill, base

ammonium cation: NIL, charge +1.

ammonium hydroxide: NHLOH

amphoteric oxide: Exhibits both sold and basic

properties

named Assessing process until \$500 masses of C-12.

arrion; ton with a regutive charge

anode: Electrole that supports existation

antibonding MO: MO is less stable than separate AOs.

AO: Abbreviation for atomic orbital

augt Asparents "dissolved in water"

Art: Symbol for the element argon, Z = 18.

aromatic: Organic compound with a benome ring

Arrhenius acid: Produces hydronium ion in water

Arrhenius basec Yields hydroxide ion in water

Arrhenius equation for rate constant (k):

 $k = Ae^{-Ea/RT}$

R. - activation energy

arsenide anion: Av., charge «It

As: Symbol for the element assenic, Z = 33.

All: Symbol for the element astation, Z = 85

aten: Symbol for pressure in "atmospheres"

atom: Fundamental unit of all matter

atomic mass number (A): Total number of protons and neutrons.

atomic number (Z): Number of protom in the nucleus.

atomic orbital: Wave motion of electrons in atoms atomic radius: Impirical measure of atom size

atomic weight: Weighted average of natural isotopes of an observed

Auc Symbol for the element gold, Z = 79

Aufbau principle: Guides the filling of electronic and solve the self that educations to

Avogadro's law; Volume is proportional to gas moles at fixed pressure and temperature

Avogadro's number: N_s = 6.022 × 10th met⁻¹

B: Symbol for the element boron, Z = 5

Ba: Symbol for the element barium, Z = 56

balanced equation: Each side of the equation has the same number of atoms of each element and charge

balanced redox: Electron loss = electron gain

barium cation: Bar, charge +2.

base definition: Sor Arrhenius base, Breested-

Lovery base, Lowis base

base ionization constant (K.): Ligalibrium constant for home dissociation.

basic solution; plif over 7

battery: Source of electrical power; galvanic cell-

Bo: Sambol for the element beryllium, Z = 4benzene: C.H.: ereanic solvent. Benzene

aromatic ring of 6 carbon atoms beryllium eation: No., charge

both (80: Greek letter that denotes radioactive particle and a number of scientific variables

beta particle: Enemetic electron from the mucleon; charge -1.

Bh: Symbol for the element bohrium, Z = 107

Bit: Symbol for the element binemath, Z = 83blearbonate anion: HCOs, charge -1

bimolecular oxygen: O:

binary compound: Fermed from 2 elements

bisulfate anion: HSO/, charge -1

Bk: Symbol for the element berkelium, Z = 97

bleach: Chemical that is a strong exidizing agent

boiling: Liquid -- gas at the boiling point

boiling point elevation: A solution has a higher

bodling point than a paint solvent

boiling point (T_s): Liquid-gas equilibrium, P = 1 atm

boiting point of water; 100°C

bond energy: Energy held by a chemical bond

bond length: Distance between 2 atoms in a bond

bond order for Lewis model: it of bonds divided by If of bounded atomics

bond order for MO treatment: # of filled bonding MOs minus # of filled antibonding MOs

bonding electrons: Form chanical bonds in a

bonding MO: MO is more stable than AOs

Biography barry PW

constant for gas, fixed T and term

By: Sumbot for the element becoming, 2 = 35

bromide anion: fit.

charges -1 Bronsted-Lowery acid:

Proton donor Bronsted-Lowry base: Proton acceptor

buffer: Composed of weak acid and weak base; serves

to keep pHI constant

cc Symbol for penti-, the SI portix for V_{mi} or 10.7

Biowte's Law

200 400 400 800

Carboxylie Acid

R-C-OH

et: Symbol for the speed of light.

*C: Celsius temperature scale

C: Sconbol for the element curbon, Z = 6.

Culti/Ost Acettic acid.

Cat Symbol for the element calcium, Z = 20

calcium cation: Ca*, charge +2

exclusions theoriday Culti-

calcium hydroxide: CarOffle catories linglish unit of heat energy

calorimetry: Study of the release or absorption of heat.

carbide anion; C*, charge -4

carbohydrate: Organic compound; carbon bonded to neveral -OH groups (sugar and stanch).

carbon dioxide: COcarbon monoxide: CO

carbonate anion: COV, charge -2.

carbonic acid: IE-CO.

carbonyl group: CO, ligand complex carboxyl group; >CO banded to 2 groups carboxyllic acid: Organic acid: R-COOH

catalyst: Accelerates a reaction but is not consumed in the reaction; lowers the activation

COUNTY OF

cathode: Electrode that supports reduction.

eathers less with a positive charge

Cd: Symbol for the element cadmium, Z = 48.

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Resources in Education .1998 Teaching Science for Understanding James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities **Blended Learning** Anthony G. Picciano, Charles D. Dziuban, Charles R. Graham, Patsy D. Moskal, 2021-09-28 Blended Learning Research Perspectives Volume 3 offers new insights into the state of blended learning an instructional modality that combines face to face and digitally mediated experiences Education has recently seen remarkable advances in instructional technologies such as adaptive and personalized instruction virtual learning environments gaming analytics and big data software This book examines how these and other evolving tools are fueling advances in our schools colleges and universities Original scholarship from education s top thinkers will prepare researchers and learning designers to tackle major issues relating to learning effectiveness diversity economies of scale and beyond Resources for Teaching Middle School Science Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-03-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter lists about 600 science

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