

Two isotopes of potassium are listed below.



1. State one similarity between potassium-37 and potassium-42 *in terms of subatomic particles.*

potassium-37 and potassium-42 have the same number of protons

2. *In terms of subatomic particles, state one difference.*

potassium-37 and potassium-42 have different numbers of neutrons

3. Compare  ${}^{37}\text{K}$  and  ${}^{42}\text{K}$  *in terms of radioactive decay.*

${}^{37}\text{K}$  and  ${}^{42}\text{K}$  have different decay modes

-or-

${}^{37}\text{K}$  is positron decay and  ${}^{42}\text{K}$  is beta decay

4. *In terms of half life, how are  ${}^{37}\text{K}$  and  ${}^{42}\text{K}$  different?*

${}^{37}\text{K}$  and  ${}^{42}\text{K}$  have different half lives

-or-

${}^{37}\text{K}$  has a half life of 1.23 seconds and  ${}^{42}\text{K}$  has a half life of 12.4 hrs

---

Given the following equilibrium reaction:



5. *In terms of the collision theory, state why adding extra  $\text{N}_2$  produces more  $\text{NH}_3$*   
more  $\text{N}_2$  molecules means more molecules present to potentially collide, thus producing more  $\text{NH}_3$

6. *In terms of LeChatelier's principle, state why adding extra  $\text{N}_2$  produces more  $\text{NH}_3$*   
adding  $\text{N}_2$  makes the equilibrium reaction shift to relieve this stress, thus producing more  $\text{NH}_3$

---

7. *In terms of ground state, excited state, and energy transitions, explain how a bright line spectrum is produced.*

When an electron returns to the ground state from the excited state, energy is released, thus producing a visible bright line spectrum

(or some variation thereunto)

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# Chemistry Regent January 2014 Answer Key

**WJ Hussar**



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