

# Experiment 29 Prelaboratory Assignment

## Bleach Analysis

Date \_\_\_\_\_ Lab Sec. \_\_\_\_\_ Name \_\_\_\_\_ Desk No. \_\_\_\_\_

- Potassium iodate,  $\text{KIO}_3$ , and the bleach's hypochlorite ion,  $\text{ClO}^-$ , are oxidizing agents. What do they oxidize in this experiment?
  - Sodium thiosulfate,  $\text{Na}_2\text{S}_2\text{O}_3$ , is a reducing agent. What does it reduce in this experiment?
- Write the balanced net-ionic equation for the standardization of the sodium thiosulfate solution by combining equations 29.7 and 29.4.
- Experimental Procedure, Part A.1. Describe the procedure for the preparation of 100 mL of a 0.010 M  $\text{KIO}_3$  solution, starting with solid  $\text{KIO}_3$  (molar mass = 214.02 g/mol).
  - Experimental Procedure, Part B.1. Describe the procedure for the preparation of 250 mL of 0.10 M  $\text{Na}_2\text{S}_2\text{O}_3$  starting with solid  $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$  (molar mass = 248.2 g/mol).
- Experimental Procedure, Part B.3. Explain why a minimum, and not a maximum, mass of ~2 g of KI is critical for the standardization of the sodium thiosulfate solution.
- Write the balanced net-ionic equation for the analysis of the "available chlorine" in a sample using sodium thiosulfate as the reducing agent by combining equations 29.2, 29.3, and 29.4.



# Chemistry Pre Lab Manual Answers

**Didier Musso**



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## **Table of Contents Chemistry Pre Lab Manual Answers**

1. Understanding the eBook Chemistry Pre Lab Manual Answers
  - The Rise of Digital Reading Chemistry Pre Lab Manual Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Pre Lab Manual Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Pre Lab Manual Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Pre Lab Manual Answers
  - Personalized Recommendations
  - Chemistry Pre Lab Manual Answers User Reviews and Ratings
  - Chemistry Pre Lab Manual Answers and Bestseller Lists
5. Accessing Chemistry Pre Lab Manual Answers Free and Paid eBooks
  - Chemistry Pre Lab Manual Answers Public Domain eBooks
  - Chemistry Pre Lab Manual Answers eBook Subscription Services
  - Chemistry Pre Lab Manual Answers Budget-Friendly Options
6. Navigating Chemistry Pre Lab Manual Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Chemistry Pre Lab Manual Answers Compatibility with Devices
  - Chemistry Pre Lab Manual Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chemistry Pre Lab Manual Answers
    - Highlighting and Note-Taking Chemistry Pre Lab Manual Answers
    - Interactive Elements Chemistry Pre Lab Manual Answers
  8. Staying Engaged with Chemistry Pre Lab Manual Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chemistry Pre Lab Manual Answers
  9. Balancing eBooks and Physical Books Chemistry Pre Lab Manual Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chemistry Pre Lab Manual Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chemistry Pre Lab Manual Answers
    - Setting Reading Goals Chemistry Pre Lab Manual Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chemistry Pre Lab Manual Answers
    - Fact-Checking eBook Content of Chemistry Pre Lab Manual Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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