

CHAPTER 17 STUDY GUIDE

Chemical Equilibrium

Section 17.1 Equilibrium: A State of Dynamic Balance

In your textbook, read about chemical equilibrium.

Complete each statement.

1. When a reaction results in almost complete conversion of reactants to products, chemists say the reaction goes to _____.
2. A reaction that can occur in both the forward and the reverse directions is called a(n) _____.
3. _____ is a state in which the forward and reverse reactions balance each other because they take place at equal rates.
4. At equilibrium, the concentrations of reactants and products are _____ but that does not mean that the amounts or concentrations are _____.
5. Equilibrium is a state of _____, not one of _____.

In your textbook, read about equilibrium expressions and constants.

For each statement below, write *true* or *false*.

- _____ 6. The law of chemical equilibrium states that at a given pressure, a chemical system may reach a state in which a particular ratio of reactant to product concentrations has a constant value.
- _____ 7. The equation $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$ is an example of a homogeneous equilibrium.
- _____ 8. If an equilibrium constant has a value less than one, the reactants are favored at equilibrium.
- _____ 9. The value for K_{eq} is constant only at a specific volume.
- _____ 10. If the equilibrium constant for a reaction at 300 K is 49.7, the concentration of the reactants will be greater than the concentration of the products.
- _____ 11. A heterogeneous equilibrium means that reactants and products are present in more than one state.
- _____ 12. The product of the forward chemical reaction is HI, for the equilibrium expression:

$$K_{\text{eq}} = \frac{[\text{HI}]^2}{[\text{H}_2][\text{I}_2]}$$

Chemistry Matter And Change Study Guide Answers

Jill Meryl Levy



Chemistry Matter And Change Study Guide Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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Glencoe Chemistry Glencoe/McGraw-Hill, 2002

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 - Chapter 3: Chemistry Matter And Change Study Guide Answers in Everyday Life
 - Chapter 4: Chemistry Matter And Change Study Guide Answers in Specific Contexts
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