

Stoichiometry and The Mole – Review Guide

Remember:

- 1 mole = 6.02×10^{23} atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



- 1) The process used to predict the amount of a product that will be formed is stoichiometry
- 2) If you know the mass of a substance, you can find the number of moles by dividing it by the...
molar mass of substance
- 3) If you know the number of particles (atoms) of a substance, you can find the number of moles by dividing it by...
Avogadro's Number
- 4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A → Moles A → Moles B → Grams B
- 5) The essential stoichiometry problem needed to convert from one chemical to another is:
Mole to Mole
- 6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?
Chemical B
- 7) What is a mole? A unit of measurement used to count atoms/molecules.
- 8) What is Avogadro's Number? 6.02×10^{23}
- 9) What does Avogadro's number tell us? The number of atoms or molecules in a mole.

Chemistry Matter And Change Stoichiometry Study Guide

Karin Nielsen-Saines



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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