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CHAPTER 2 THE COMPONENTS OF MATTER

FOLLOW-UP PROBLEMS

- 2.1A **Plan:** An element has only one kind of atom; a compound is composed of at least two kinds of atoms. A mixture consists of two or more substances mixed together in the same container.
Solution:
(a) There is only one type of atom (blue) present, so this is an **element**.
(b) Two different atoms (brown and green) appear in a fixed ratio of 1/1, so this is a **compound**.
(c) These molecules consist of one type of atom (orange), so this is an **element**.
- 2.1B **Plan:** An element has only one kind of atom; a compound is composed of at least two kinds of atoms.
Solution:
The circle on the left contains molecules with either only orange atoms or only blue atoms. This is a **mixture of two different elements**. In the circle on the right, the molecules are composed of one orange atom and one blue atom so this is a **compound**.
- 2.2A **Plan:** Use the mass fraction of uranium in pitchblende (from Sample Problem 2.2) to find the mass of pitchblende that contains 2.3 t of uranium. Subtract the amount of uranium from that amount of pitchblende to obtain the mass of oxygen in that amount of pitchblende. Find the mass fraction of oxygen in pitchblende and multiply the amount of pitchblende by the mass fraction of oxygen to determine the mass of oxygen in the sample.
Solution:
Mass (t) of pitchblende = 2.3 t uranium $\times \frac{84.2 \text{ t pitchblende}}{71.4 \text{ t uranium}} = 2.7123 \text{ t pitchblende}$
Mass (t) of oxygen in 84.2 t of pitchblende = 84.2 t pitchblende – 71.4 t uranium = 12.8 t oxygen
Mass (t) of oxygen = 2.7123 t pitchblende $\times \frac{12.8 \text{ t oxygen}}{84.2 \text{ t pitchblende}} = 0.4129 \text{ t oxygen}$
- 2.2B **Plan:** Subtract the amount of silver from the amount of silver bromide to find the mass of bromine in 26.8 g of silver bromide. Use the mass fraction of silver in silver bromide to find the mass of silver in 3.57 g of silver bromide. Use the mass fraction of bromine in silver bromide to find the mass of bromine in 3.57 g of silver bromide.
Solution:
Mass (g) of bromine in 26.8 g silver bromide = 26.8 g silver bromide – 15.4 g silver = 11.4 g bromine
Mass (g) of silver in 3.57 g silver bromide = 3.57 g silver bromide $\times \frac{15.4 \text{ g silver}}{26.8 \text{ g silver bromide}} = 2.05 \text{ g silver}$

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Table of Contents Chemistry Matter And Change Assessment Solution Manual

1. Understanding the eBook Chemistry Matter And Change Assessment Solution Manual
 - The Rise of Digital Reading Chemistry Matter And Change Assessment Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Matter And Change Assessment Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Matter And Change Assessment Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Matter And Change Assessment Solution Manual
 - Personalized Recommendations
 - Chemistry Matter And Change Assessment Solution Manual User Reviews and Ratings

- Chemistry Matter And Change Assessment Solution Manual and Bestseller Lists
- 5. Accessing Chemistry Matter And Change Assessment Solution Manual Free and Paid eBooks
 - Chemistry Matter And Change Assessment Solution Manual Public Domain eBooks
 - Chemistry Matter And Change Assessment Solution Manual eBook Subscription Services
 - Chemistry Matter And Change Assessment Solution Manual Budget-Friendly Options
- 6. Navigating Chemistry Matter And Change Assessment Solution Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Matter And Change Assessment Solution Manual Compatibility with Devices
 - Chemistry Matter And Change Assessment Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Matter And Change Assessment Solution Manual
 - Highlighting and Note-Taking Chemistry Matter And Change Assessment Solution Manual
 - Interactive Elements Chemistry Matter And Change Assessment Solution Manual
- 8. Staying Engaged with Chemistry Matter And Change Assessment Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Matter And Change Assessment Solution Manual
- 9. Balancing eBooks and Physical Books Chemistry Matter And Change Assessment Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Matter And Change Assessment Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Matter And Change Assessment Solution Manual
 - Setting Reading Goals Chemistry Matter And Change Assessment Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Matter And Change Assessment Solution Manual
 - Fact-Checking eBook Content of Chemistry Matter And Change Assessment Solution Manual
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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