

General Chemistry II Chapter 17 Study Guide

Organic acids- carboxylic acids (R-COOH):

- Carboxylic acids (R-COOH): a weak organic acid; contains the carboxyl group (-COOH); never fully dissociates in water; an equilibrium process; ex. Formic acid, acetic acid, butyric acid, benzoic acid
 - R= rest of the molecule
- Strength of carboxylic acids: depends on the R group; more electronegative Cl in R group shifts more electron density towards the R group and helps stabilize the structure after the H^+ ion is removed; thus, chloroacetic acid is the stronger acid in comparison to acetic acid (CH_3COOH)

Organic bases- amines:

- Amines: organic bases; contains an N atom with a lone pair (ex. $-NH$, $-NH_2$); ex. Methyl amine, dimethyl amine, trimethyl amine; are weak bases:
 - Only N with lone pair can accept a proton
 - Equilibrium dissociation process
 - NH_4^+ is an acid and can lose one H^+ to be neutral

Organic acids and bases (things to keep in mind):

- Acid: H is the only acidic proton; ex. CH_3COOH becomes CH_3COO^-
- Amine: the proton H^+ is bonded to N via the lone pair of N; ex. CH_3NH_2 becomes $CH_3NH_3^+$
- How about carbon?
 - Protons on C are not acidic; C cannot accept extra protons

Autoionization of water:

- Autoionization: in pure water, one water molecule can donate a proton to another water molecule



At 25 °C, $[H_3O^+] = [OH^-] = 1.0 \times 10^{-7} M$

$$K_w = [H_3O^+] [OH^-] = 1.0 \times 10^{-14}$$

K_w : ion-product constant

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Jianjun Gao



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