AP Chem Chapter 16 Practice Quiz:

- 1. Which statement is FALSE?
- a) If a reaction is thermodynamically spontaneous, it may occur rapidly.
- b) A fast reaction may be thermodynamically spontaneous.
- c) If a reaction is thermodynamically spontaneous, it may occur slowly.
- d) If a reaction is thermodynamically spontaneous, it must have a low activation energy.
- e) Rate of reaction is a kinetic quantity rather than a thermodynamic quantity.
- 2. Which of the following is a kinetics concept?
- a) firee emergy
- b) enthalpy
- c) spontaneity
- d) reaction mechanism
- e) entropy
- 3. A cutalyst:
- a) is used up in a chemical reaction
- b) changes the value of A G of the reaction
- c) is always a solid
- d) does not influence the reaction in any way
- e) changes the activation energy of the reaction
- 4. What is the entropy change of the reaction below at 298K and 1 atm pressure?

	Note:	3H ₂ (g)>	2NH ₂ (g)
S(U/mol x K)	191.5	1.30.6	192.3

- a) -198.7 J/K
- b) 76.32 J/K
- c) -129.7 J/K
- db 3003.2 J/80
- e) 384.7 J/K
- If the entropy change for the reaction below at 298 K and 1 atm pressure is 137 J/K, and S=205 J/mol x K for O₂(g), what is S for O₃(g)?

$$2O_3 ---> 3O_3$$

- a) 364 J/mol x K
- b) 478 J/mol x K
- e) 239 Jimel x K.
- d) -117 J/mol x K.
- e.) -59 J/mod.x K
- 6. The heat of vaporization of methanol, CH₂OH, is 35.20 kJ/mol. Its boiling point is 64.6 degrees Celsius. What is the change in entropy for the vaporization of methanol?
- a) -17.0 J/mol x K.
- b) 3.25 J/mol x K.
- c) 17.0 J/mel x K.
- d) 104 J/mol x K.
- c) 543 J/mol x K.
- 7. Which one on the following reactions has a positive entropy change?
- a) H₂O(g) ----> H₂O(l)
- b) $BF_3(g) + NH_3(g) \longrightarrow F_3BNH_3(g)$
- c) $2SO_2(g) + O_2(g) ---> 2SO_3(g)$
- d) $N_2(g) + 3H_2(g) ---> 2NH_2(g)$
- e) $2NH_4NO_3(s) \longrightarrow 2N_2(g) + 4H_2O(g) + O_2(g)$

Chemistry Ch 16 Test

James Girard

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