

12: Stoichiometry

Key Stoichiometry Terms

- Stoichiometry:** Using the mole ratio in the balanced equation and information about one compound to find information about another in the reaction.
- Dimensional Analysis:** Method of converting units by multiplying by ratio of equalities.
- Molar Mass:** Sum of all the atomic masses (from the periodic table) in the compound.
- Solute:** Substance being dissolved in a homogeneous mixture (solution).
- Concentration:** Measure of how much solute is dissolved in how much solution.
- Molarity (M):** A concentration unit in moles per unit liter.
- Standard Temperature and Pressure (STP):** 1 atm (or 101.3 kPa) and 273 K (0°).
- Molar Volume of a Gas:** 1 mole of any gas at STP is 22.4 liters.
- Limiting Reactant:** Reactant that stops the reaction by running out first.
- Actual Yield:** The amount actually produced in the lab.
- Theoretical Yield:** Amount that should be theoretically produced based on stoichiometric calculations.
- Percent Yield:** Compares the actual yield to the theoretical yield (their ratio with %).

Chemical Equations

- Coefficients give the mole ratios.

Example:

$2 \text{H}_2 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow 2 \text{H}_2\text{O} (\text{l})$
2 moles of H_2 react with 1 mole of O_2 and produces 2 moles of H_2O .

Dimensional Analysis in Stoichiometry

Equalities used during dimensional analysis for stoichiometry:

- Mole ratio in balanced equation:** Use to convert between moles of different compounds in the balanced equation.
- Molar mass:** Used to convert between grams and moles.
- Concentration:** Used to convert between moles and liters of a solution.
- Molar volume of a gas:** Used to convert between moles and liters of a gas at STP.

KUDOS for Stoichiometric Problems

Method to solve word problems:

- K = Known.** Identify all the known information (given quantities, balanced equation, etc.).
- U = Unknown.** What quantity and what compound does the problem ask for?
- D = Definition.** Identify the equalities that will be needed during dimensional analysis to convert from the known to the unknown.
- O = Output.** Perform the dimensional analysis.
- S = Substantiation.** Check for reasonableness, units and correct significant figures.

Molarity

Molarity - the concentration measure by the moles of solute per liter of solution.

$$\text{Molarity} = \frac{\text{moles solute}}{\text{L solution}}$$

Limiting Reactants

- Once a reactant has run out, the reaction will stop.
 - Do stoichiometry for each given reactant quantity to the same product each time. Choose the calculation that gives the **smallest amount** of product.
 - The reactant that produced the smallest amount of product is the **limiting reactant**.
 - Steps to determine the limiting reactant: grams reactant \rightarrow moles reactant \rightarrow moles product \rightarrow grams product. Alternatively, use the m/c method below.
- Simple Limiting Reagent Finder Mnemonic:** Find the moles of each reactant and divide the moles of each reactant by its coefficient. The reactant with the smallest number is the limiting reagent = "Smallest m/c is limited".

Percent Yield

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

Example:

If stoichiometry calculations tell you that a reaction should produce 10.5 g, but when you perform the lab you produce 8.7 g, what was the percent yield?

$$\% \text{ yield} = \frac{8.7 \text{ g}}{10.5 \text{ g}} \times 100 = 82.9\%$$

Stoichiometry Examples

All examples for the following reaction:



Mole-Mass example:

0.75 moles HCl react. How many grams Mg are needed?

$$\frac{0.75 \text{ mole HCl}}{2 \text{ mole HCl}} \times \frac{1 \text{ mole Mg}}{1 \text{ mole HCl}} \times \frac{24.31 \text{ g Mg}}{1 \text{ mole Mg}} = 9.1 \text{ g Mg}$$

Mass-Mass example:

If 2.5 g Mg react, how many grams MgCl_2 are produced?

$$\frac{2.5 \text{ g-Mg}}{24.31 \text{ g-Mg}} \times \frac{1 \text{ mole-Mg}}{1 \text{ mole-Mg}} \times \frac{1 \text{ mole-MgCl}_2}{1 \text{ mole-Mg}} \times \frac{95.21 \text{ g-MgCl}_2}{1 \text{ mole-MgCl}_2} = 9.8 \text{ g MgCl}_2$$

Mass-Liter example (solution):

How many liters of 1.7M HCl are needed to react with 2.5 g Mg?

$$\frac{2.5 \text{ g-Mg}}{24.31 \text{ g-Mg}} \times \frac{1 \text{ mole-Mg}}{1 \text{ mole-Mg}} \times \frac{2 \text{ mole-HCl}}{1 \text{ mole-Mg}} \times \frac{1 \text{ L HCl}}{1.7 \text{ mole HCl}} = 0.12 \text{ L HCl}$$

Mass-Liter example (gas):

How many liters of H_2 are produced when 3.5 g Mg react?

$$\frac{3.5 \text{ g-Mg}}{24.31 \text{ g-Mg}} \times \frac{1 \text{ mole-Mg}}{1 \text{ mole-Mg}} \times \frac{1 \text{ mole-H}_2}{1 \text{ mole-Mg}} \times \frac{22.4 \text{ L H}_2}{1 \text{ mole H}_2} = 3.2 \text{ L H}_2$$

Limiting reactant:

If 0.25 mole HCl reacts with 0.55 mole Mg, how many grams of MgCl_2 are produced and what was the limiting reactant?



$$\frac{0.25 \text{ mole HCl}}{2 \text{ mole HCl}} \times \frac{1 \text{ mole-MgCl}_2}{1 \text{ mole-Mg}} \times \frac{95.21 \text{ g-MgCl}_2}{1 \text{ mole-MgCl}_2} = 12 \text{ g MgCl}_2$$

(less)

$$\frac{0.55 \text{ mole-Mg}}{1 \text{ mole-Mg}} \times \frac{1 \text{ mole-MgCl}_2}{1 \text{ mole-Mg}} \times \frac{95.21 \text{ g-MgCl}_2}{1 \text{ mole-MgCl}_2} = 52 \text{ g MgCl}_2$$

12 g MgCl_2 will be produced and HCl is the limiting reactant.

How to Use This Cheat Sheet: These are the keys related to this topic. Try to read through it carefully twice then write it out from memory on a blank sheet of paper. Review it again before the exams.

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Acing the ACS James Coy-Dibley, 2017-09-05 Study guide for first year general chemistry ACS Exam

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Eubanks,1998 *Barron's AP Chemistry* Neil D. Jespersen,Pamela Kerrigan,2016-03-01 Extensive test preparation for the AP Chemistry exam includes Six practice AP exams three diagnostic tests and three full length practice exams that are aligned with the upcoming new AP Chemistry exam All questions answered and explained A comprehensive subject review covering the structure of matter chemical bonding states of matter physical chemistry chemical reactions and all other test topics Study tips and test taking strategies This review book can be purchased alone or with an optional CD ROM that presents two additional full length practice tests with automatic scoring and fully explained answers GACE Chemistry Study Guide GACE Chemistry Exam Prep Team,2018-02-12 Imagine a study guide actually designed for teachers Because we know you ve got a busy life we ve developed a study guide that isn t like other certification materials out there With our NEW GACE Chemistry Study Guide Test Prep and Practice Questions for the GACE Chemistry Exam you get a swift but full review of everything tested on your certification exam FREE online resources are also included with your study guide Imagine having FREE practice questions digital flash cards study cheat sheets and 35 test tips available anytime anywhere on your cell phone or tablet Our resources will give you the push you need to pass your test the first time Because GACE Chemistry Study Guide tells you everything you need to know and nothing you don t you ll have more time to concentrate on what s important in your life Furthermore GACE Chemistry Study Guide gives you the opportunity to test your knowledge and assess your skills with practice tests and answers based on the actual exam Because teaching is more than a test we give you JUST what you need to succeed GACE Chemistry Study Guide Test Prep and Practice Questions for the GACE Chemistry Exam will provide you with a detailed overview of the GACE Chemistry exam so you know exactly what to expect on test day We ll take you through all the concepts covered on the test and give you the opportunity to test your knowledge with practice questions Even if it s been a while since you last took a major test don t worry we ll make sure you re more than ready Cirrus Test Prep s GACE Chemistry Study Guide Test Prep and Practice Questions for the GACE Chemistry Exam includes a comprehensive REVIEW of Basic Principles of Matter Atomic and Nuclear Structure Bonding Naming Compounds Chemical Reactions Thermodynamics Solutions and Acid Base Chemistry Scientific Inquiry and Procedures as well as a FULL GACE Chemistry practice test About Cirrus Test Prep Developed by experienced current and former educators Cirrus Test Prep s study materials help future educators gain the skills and knowledge needed to successfully pass their state level teacher certification exams and enter the classroom Each Cirrus Test Prep study guide includes a detailed summary of the test s format content and scoring an overview of the content knowledge required to pass the exam worked through sample questions with answers and explanations full length practice tests including answer explanations and unique test taking strategies with highlighted key concepts Cirrus Test Prep s study materials ensure that new educators feel prepared on test day and beyond **Chemists' Guide to Effective Teaching** Norbert J. Pienta,Melanie M. Cooper,Thomas J. Greenbowe,2005 For courses in Methods of Teaching Chemistry Useful for new professors chemical educators or students

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- Chapter 5: Conclusion

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