

Chemistry 30

Unit 5: Acids & Bases

Assignment 2 – K_a , K_b , K_w and pH

39

- 3 1. Given the following balanced ionization reactions for the following weak acids and bases, write the K_a or K_b expressions for each.



$$K_a = \frac{[\text{H}^+][\text{C}_6\text{H}_7\text{O}_5^-]}{[\text{HC}_6\text{H}_7\text{O}_5]}$$



$$K_a = \frac{[\text{H}^+][\text{H}_2\text{BO}_3^-]}{[\text{H}_3\text{BO}_3]}$$



$$K_b = \frac{[\text{CH}_3\text{NH}_3^+][\text{OH}^-]}{[\text{CH}_3\text{NH}_2]}$$

- 2 2. Calculate $[\text{OH}^-]$ in a solution containing 100.0 g of potassium hydroxide in 2.50 L solution. Potassium hydroxide is a strong base.

The molar mass of KOH is $56.1 \text{ g} \cdot \text{mol}^{-1}$

$$[\text{KOH}] = \frac{\text{mol}}{56.1\text{g}} \times \frac{100.0\text{g}}{1} \times \frac{1}{2.50\text{L}} = \frac{0.713\text{mol}}{\text{L}} = 0.713\text{M}$$



Since KOH is a strong base, based on the balanced equation $[\text{OH}^-] = [\text{KOH}] = 0.713 \text{ M}$

- 2 3. A solution is prepared in which 0.600 mole of hydrogen chloride is dissolved in enough water to make 5.80 L. Calculate the concentration of hydrogen ions in this solution.

$$[\text{HCl}] = \frac{0.600\text{mol}}{5.80\text{L}} = \frac{0.103\text{mol}}{\text{L}} = 0.103\text{M}$$



Since HCl is a strong acid, based on the balanced equation $[\text{H}^+] = [\text{HCl}] = 0.103 \text{ M}$

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