

**Question 1. Calculate the molecular mass of the following:**

**(i)  $\text{H}_2\text{O}$ (ii)  $\text{CO}_2$ (iii)  $\text{CH}_4$**

**Answer:** (i) Molecular mass of  $\text{H}_2\text{O} = 2(1.008 \text{ amu}) + 16.00 \text{ amu} = 18.016 \text{ amu}$

(ii) Molecular mass of  $\text{CO}_2 = 12.01 \text{ amu} + 2 \times 16.00 \text{ amu} = 44.01 \text{ amu}$

(iii) Molecular mass of  $\text{CH}_4 = 12.01 \text{ amu} + 4 (1.008 \text{ amu}) = 16.042 \text{ amu}$

**Question 2. Calculate the mass percent of different elements present in sodium sulphate ( $\text{Na}_2 \text{SO}_4$ ).**

**Answer:**

$$\text{Mass \% of an element} = \frac{\text{Mass of that element in the compound}}{\text{Molar mass of the compound}} \times 100$$

Now,  $\text{Molar mass of } \text{Na}_2\text{SO}_4 = 2 (23.0) + 32.0 + 4 \times 16.0$   
 $= 142 \text{ g mol}^{-1},$

$$\begin{aligned}\text{Mass percent of sodium} &= \frac{46}{142} \times 100 \\ &= 32.39 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percent of sulphur} &= \frac{32}{142} \times 100 \\ &= 22.54 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percent of oxygen} &= \frac{64}{142} \times 100 \\ &= 45.07 \%\end{aligned}$$

**Question 3. Determine the empirical formula of an oxide of Iron which has 69.9 % iron and 30.1 % dioxygen by mass.**

# Chemistry 11th Numerical Solution

**Yijin Wang**



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### **Table of Contents Chemistry 11th Numerical Solution**

1. Understanding the eBook Chemistry 11th Numerical Solution
  - The Rise of Digital Reading Chemistry 11th Numerical Solution
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 11th Numerical Solution
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry 11th Numerical Solution
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 11th Numerical Solution
  - Personalized Recommendations
  - Chemistry 11th Numerical Solution User Reviews and Ratings
  - Chemistry 11th Numerical Solution and Bestseller Lists

5. Accessing Chemistry 11th Numerical Solution Free and Paid eBooks
  - Chemistry 11th Numerical Solution Public Domain eBooks
  - Chemistry 11th Numerical Solution eBook Subscription Services
  - Chemistry 11th Numerical Solution Budget-Friendly Options
6. Navigating Chemistry 11th Numerical Solution eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry 11th Numerical Solution Compatibility with Devices
  - Chemistry 11th Numerical Solution Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry 11th Numerical Solution
  - Highlighting and Note-Taking Chemistry 11th Numerical Solution
  - Interactive Elements Chemistry 11th Numerical Solution
8. Staying Engaged with Chemistry 11th Numerical Solution
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry 11th Numerical Solution
9. Balancing eBooks and Physical Books Chemistry 11th Numerical Solution
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry 11th Numerical Solution
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 11th Numerical Solution
  - Setting Reading Goals Chemistry 11th Numerical Solution
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 11th Numerical Solution
  - Fact-Checking eBook Content of Chemistry 11th Numerical Solution
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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