

# Chemistry Formulas

Dilution

$$M_1V_1 = M_2V_2$$

Energy of a photon

$$E_{\text{photon}} = hf = hc/\lambda$$

Pressure

$$P = F/A$$

Average Kinetic Energy

$$KE_{\text{avg}} = \frac{3}{2} RT$$

Ideal Gas Law

$$PV = nRT$$

Combined Gas Law

$$\frac{P_1V_1}{N_1T_1} = \frac{P_2V_2}{N_2T_2}$$

Dalton's Law of Partial Pressures

$$P_{\text{total}} = P_A + P_B + \dots$$

Dalton's Law of Partial Pressures

$$P_A = \chi_A P_{\text{total}}$$

Graham's Law of Effusion

$$\frac{r_A}{r_B} = \sqrt{\frac{M_B}{M_A}}$$

Henry's Law

$$P_A = k_H[A]$$

Freezing Point Depression

$$\Delta T_f = -iK_f M$$

Boiling Point Elevation

$$\Delta T_b = iK_b m$$

Vapor Pressure Depression (Raoult's Law)

$$P_{\text{soln}} = \chi_{\text{solvent}} P_{\text{solvent}}^0$$

Osmotic Pressure ( $\pi$ )

$$\pi = iMRT$$

Arrhenius Equation

$$k = Ae^{-E_a/RT}$$

Equilibrium Constant Expressions

$$K_c = \frac{[\text{products}]}{[\text{reactants}]}$$

Equilibrium Constant Expressions

$$K_{eq} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$$

Equilibrium Constant Expressions

$$K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Reaction Quotient (Q)

$$Q = \frac{[\text{products}]}{[\text{reactants}]}$$

Solubility Product Constant ( $K_{sp}$ )

$$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$$

pH & pOH

$$pH = -\log[H^+]$$

pH & pOH

$$pOH = -\log[OH^-]$$

pH & pOH

$$pH + pOH = 14$$

$pK_a$  &  $pK_b$

$pK_a$  &  $pK_b$

$pK_a$  &  $pK_b$

Neutralization Reaction

Buffers

Enthalpy of Formation

First Law of Thermodynamics

Pressure Volume Work

Calorimetry Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Entropy (S)

Bond Dissociation Energy

Gibb's Free Energy ( $\Delta G$ )

Gibb's Free Energy ( $\Delta G$ )

Gibb's Free Energy ( $\Delta G$ )

Standard Cell Potential

Standard Cell Potential

Nernst Equation

Kinetics (always 1st order)

Kinetics (always 1st order)

Kinetics (always 1st order)

Nuclear Binding Energy

$$pK_a = -\log[K_a]$$

$$pK_b = -\log[K_b]$$

$$pK_a + pK_b = 14$$

$$n_A M_A V_A = n_B M_B V_B$$

$$pH = pK_a + \log[A^-]/[HA]$$

$$\Delta H_f = \sum n \Delta H_f^{\circ}(\text{product}) - \sum n \Delta H_f^{\circ}(\text{reactants})$$

$$\Delta E = q + w$$

$$w = -P\Delta V$$

$$q = -C_{\text{calorimeter}} \Delta T$$

$$q = mc\Delta T$$

$$q = m\Delta H_{\text{fusion}}$$

$$q = m\Delta H_{\text{vaporization}}$$

$$\Delta S = \ln S_{\text{products}} - \ln S_{\text{reactants}}$$

$$\Delta H = \sum \Delta H_{\text{reactants}} - \sum \Delta H_{\text{products}}$$

$$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$$

$$\Delta G = \Delta G^{\circ} + RT \ln Q$$

$$\Delta G^{\circ} = -RT^{\circ} \ln(K_{eq})$$

$$E^{\circ} = E^{\circ}_{\text{reduction}} + E^{\circ}_{\text{oxidation}}$$

$$E^{\circ} = E^{\circ}_{\text{cathode}} + E^{\circ}_{\text{anode}}$$

$$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$$

$$N = N_0 e^{-kt}$$

$$\ln N = \ln N_0 - kt$$

$$t_{1/2} = \frac{0.693}{k}$$

$$E = \Delta mc^2$$

# Chemical Names And Formulas Chemistry Study Guide

**J Rink**



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