

## 16 • Chemical Equilibrium

1. Consider the reaction system,



The equilibrium constant expression is

- a)  $\frac{[\text{CoO}][\text{H}_2]}{[\text{Co}][\text{H}_2\text{O}]}$       d)  $\frac{[\text{H}_2]}{[\text{H}_2\text{O}]}$   
 b)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{CoO}][\text{H}_2]}$       e)  $\frac{[\text{H}_2\text{O}]}{[\text{H}_2]}$   
 c)  $\frac{[\text{Co}][\text{H}_2\text{O}]}{[\text{H}_2]}$

2. Given the equilibrium,

if this equilibrium is established by beginning with equal number of moles of  $\text{SO}_2$  and  $\text{O}_2$  in a 1.0 Liter bulb, thenthe following *must* be true at equilibrium:

- a)  $[\text{SO}_2] = [\text{SO}_3]$       d)  $[\text{SO}_2] < [\text{O}_2]$   
 b)  $2[\text{SO}_2] = 2[\text{SO}_3]$       e)  $[\text{SO}_2] > [\text{O}_2]$   
 c)  $[\text{SO}_2] = [\text{O}_2]$

Questions 3 &amp; 4 refer to the following:

At a given temperature, 0.300 mole  $\text{NO}$ , 0.200 mol  $\text{Cl}_2$  and 0.500 mol  $\text{ClNO}$  were placed in a 25.0 Liter container. The following equilibrium is established:  $2\text{ClNO(g)} \rightleftharpoons 2\text{NO(g)} + \text{Cl}_2\text{(g)}$

3. At equilibrium, 0.600 mol of
- $\text{ClNO}$
- was present. The number of
- moles*
- of
- $\text{Cl}_2$
- present at equilibrium is

- a) 0.050      d) 0.200  
 b) 0.100      e) 0.250  
 c) 0.150

## P R A C T I C E T E S T

4. The equilibrium constant,
- $K_c$
- , is:

- a)  $4.45 \times 10^{-4}$       d) 0.167  
 b)  $6.67 \times 10^{-4}$       e) 1500  
 c) 0.111

5. At
- $985^\circ\text{C}$
- , the equilibrium constant for the reaction,



is 1.63. What is the equilibrium constant for the reverse reaction?

- a) 1.63      d) 0.613  
 b) 0.815      e) 1.00  
 c) 2.66

6. What is the relationship between
- $K_p$
- and
- $K_c$
- for the reaction,
- $2\text{HCl(g)} \rightleftharpoons \text{H}_2\text{(g)} + \text{Cl}_2\text{(g)}$
- ?

- a)  $K_p = K_c(RT)^{-1}$       d)  $K_p = K_c$   
 b)  $K_p = K_c(RT)$       e)  $K_p = K_c(2RT)$   
 c)  $K_p = K_c(RT)^2$

7. For the reaction
- $2\text{NO}_2\text{(g)} \rightleftharpoons \text{N}_2\text{O}_4\text{(g)}$
- ,
- $K_p$
- at
- $25^\circ\text{C}$
- is 7.3, when all partial pressures are expressed in atmospheres. What is
- $K_c$
- for this reaction? [
- $R=0.0821 \text{ L}\cdot\text{atm}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$
- ]

- a) 4270      d) 179  
 b) 0.0119      e) 2.06  
 c) 0.291

# Chemical Equilibrium Test

**Harvey F. Lodish**



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A. Mascetta, Mark Kernion, 2020-12-01 Note College Board has discontinued the SAT Subject Tests in the US. The tests will be available outside the US in June 2021 and then be discontinued. Barron's SAT Subject Test Chemistry with 7 Practice Tests features in-depth review of all topics on the exam and full-length practice tests in the book and online. This edition includes One full-length diagnostic test to help you assess your strengths and weaknesses. Comprehensive review of all topics on the exam including introductory chemistry, atomic structure and the periodic table, bonding, chemical formulas, gases and laws, stoichiometry, liquids, solids and phase changes, chemical reactions and thermochemistry, chemical reactions, chemical equilibrium, acids, bases and salts, oxidation-reduction, carbon and organic chemistry and the laboratory. Four full-length practice tests that reflect the actual SAT Subject Test Chemistry exam in length, question types and degree of difficulty. Two full-length online practice tests with answer explanations and automated scoring. Appendices which include the periodic table, important equation, constant and data tables and a glossary of chemistry terms. **English & General Aptitude for**

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