Chapter 18

Study Guide - Answers

- Write the expression for the equilibrium constant for each of the following reactions. 100
 - a. H₂(g) + CI₂(g) = 2HCl(g)

$$K_{eq} = \frac{[HCI]^2}{[H_2][CI_2]}$$

b.
$$N_2(g) + 3 H_2(g) = 2 NH_1(g)$$

$$K_{eq} = \frac{[NH_1]^4}{[N_2][H_2]^4}$$
c. $HCN(g) \stackrel{\text{def}}{=} H^4(aq) + CN^4(aq)$

d.
$$H_2SO_d(g) \stackrel{\text{def}}{=} H'(aq) + HSO^*(aq)$$

$$K_{eq} = \frac{-[H'][HSO^4]}{[H:SO_4]}$$

e.
$$PbF_2(s) = Pb^{21}(aq) + 2F(aq)$$

2 From the data provided below, calculate the value of the equilibrium constant for the reaction.

$$[H_2] = [CI_2] = 1.0 \times 10^{-2}; [HCI] = 1.0 \times 10^{-4}$$

$$K_{eq} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$

b.
$$N_2(g) + 3 H_2(g) = 2 NH_2(g)$$

$$[N_3] = 4.4 \times 10^{-2}$$
; $[H_3] = 1.2 \times 10^{-1}$; $[NH_3] = 3.4 \times 10^{-3}$

$$K_{sq} = \frac{[3.4 \times 10^{-3}]^2}{[4.4 \times 10^{-3}][1.2 \times 10^{-3}]^3} = 0.152$$

3. For the reaction $N_2O_4(g) = 2NO_2(g)$, $K_{eq} = 0.2$. At a particular time, the following concentrations are measured; [N₂O₄] = 2.0 M, [NO₂] = 0.2 M. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?

$$N_2O_2(g) \cong 2NO_2(g)$$

$$[N_2O_3] = 2.0 \text{ M}$$

 $[NO_2] = 0.2 \text{ M}$

$$Q = -\frac{[NO_2]^2}{[N_2O_4]}$$

$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{eq} = 0.2$$

 $Q \le K_{eq}$ so the reaction will proceed to the right (forward).

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