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Chemical Reactions Study Guide

I. Balancing Chemical Equations

Notes #5 The first step to balancing chemical reaction equations is to be able to accurately count the number of each type of element on both the reactants and products side of the arrow. This is done by multiplying the <u>coefficient</u> (big number in front of an element symbol) with the <u>subscript</u> (little number found after the element symbol). Remember: if no coefficient or subscript is written, then there is an implied, imaginary "1".

Examples: 3 Br. there are 6 bromine atoms (3 × 2 = 6)

CaO \leftarrow there is 1 calcium and 1 oxygen (in each case 1 × 1 = 1) The "1"s are implied. CH₄ + O₁ \rightarrow H₂O + CO₂ \leftarrow There are 3 total oxygens on the product (right) side (1 × 1 + 1 × 2 = 3)

Practice: How many of each type of atom is represented.

3 Cl₂ Fe 2 S₈ MgF₂

L₂ 4 Ni 2 Li₂O H₂O₂

H₂SO₄ 2 Pb(NO₂)₂

C₂H₄ + O₂ → H₂O + CO₂ Resoctants

Products _____

Notes #6 Now that we've learned to count, it's time to put our skills into action by altering the numbers of certain elements to suit our desires. In any case, our desires are to have the same number of each type of element on both the <u>products</u> and <u>reactants</u> side of the equation. We do this by <u>changing the coefficients</u>.

Example: Li + O₂ → Li₂O A lovely synthesis reaction. What do we do first?

Count the number of each element on both sides.

Reactants: 1Ll and 2 O's Products: 2 Ll's and 1 O

Add/change coefficients to balance.

1) First, I check the first element, Li (lithium). There's one on the reactants side and 2 on the

products, so I'll add a coefficient of 2 to the reactants to balance: ∠U • O₁ → U.O

 Now for O (oxygen). 2 on reactants, 1 on product so I'll put a 2 in front of Li₂O to balance: 2 Li • O₁ → 2 Li₂O

 Whoops, that affects lithium. I now have four on the products side, so I'll go back and change the "2" to a "4" on the reactants side: 4.Li + O₂ → 2 Li₂O BALANCEDIII Ø

Practice: Balance each of the following chemical reaction equations. Check your answer by counting each element on each side when you think you're done.

 $_{_{_{_{_{_{1}}}}}}$ $H_{1}+_{_{_{_{_{_{_{1}}}}}}}$ $O_{2} \rightarrow _{_{_{_{_{_{1}}}}}}$ $H_{1}O$ $_{_{_{_{_{_{1}}}}}}$ $N_{3}+_{_{_{_{_{_{1}}}}}}$ $N_{3}+_{_{_{_{_{_{1}}}}}}$ $N_{3}+_{_{_{_{_{1}}}}}$ $N_{3}+_{_{_{_{_{1}}}}}$

____NaHCO₃ → ____Na₂CO₃ + ____CO₂ + ____H₂O

_____Al + _____CuO -> _____Al₂O₂ + _____Cu

 $K_2O + \underline{\hspace{1cm}} H_2O \rightarrow \underline{\hspace{1cm}} KOH \underline{\hspace{1cm}} Mg + \underline{\hspace{1cm}} P_i \rightarrow \underline{\hspace{1cm}} Mg_2P_2$

 $C_2H_4 + C_2 \rightarrow H_2O + CO_2$

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