517. At a temperature of 1,800 K, the reactions represented by the following two equations predominate in air at 1.0 atm pressure:

$$1/2 N_2 + 1/2 O_2 \Rightarrow NO$$
 (1)

$$NO_2 = NO + 1/2 O_2$$
 (2)

Assume the following:

For air:

$$O_2 = 21 \text{ vol}\%$$
  
 $N_2 = 79 \text{ vol}\%$ 

x = NO formed at equilibrium, mol

y = NO2 formed at equilibrium, mol

Which of the following expressions for the equilibrium constant is correct for Reaction 1?

(A) 
$$\frac{x}{(0.79-x/2-y/2)^{1/2}(0.21-x/2-y)^{1/2}}$$

(B) 
$$\frac{x}{(1-y/2)(0.79-x/2-y/2)^{1/2}(0.21-x/2-y)^{1/2}}$$

(C) 
$$\frac{x(1-y/2)}{(0.79-x/2-y/2)^{1/2}(0.21-x/2-y)^{1/2}}$$

(D) 
$$\frac{x}{(0.79 - x/2 - y)^{1/2}(0.21 - x/2 - y/2)^{1/2}}$$

# **Chemical Engineering Practice Pe Exams**

Dilip K. Das, Rajaram K. Prabhudesai

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