

Chemical Reactions

Section 9.1 Reactions and Equations

In your textbook, read about evidence of chemical reactions.

For each statement, write *yes* if evidence of a chemical reaction is present. Write *no* if there is no evidence of a chemical reaction.

- | | |
|---------------|---|
| <u>yes</u> | 1. A tomato smells rotten. |
| <u>no</u> | 2. A drinking glass breaks into smaller pieces. |
| <u>no</u> | 3. A piece of ice melts. |
| <u>yes</u> | 4. Drain cleaner is mixed with water and the solution becomes warm. |
| <u>yes</u> | 5. Candle wax burns. |
| <u>yes/no</u> | 6. Molten candle wax solidifies. |
| <u>yes</u> | 7. Green leaves turn yellow and red as the seasons change. |
| <u>yes</u> | 8. Baking powder produces a gas that makes a cake rise. |

In your textbook, read about how to represent chemical reactions and how to balance chemical equations.

Use the terms below to complete the passage. Each term may be used once, more than once, or not at all.

arrow	plus sign	(s)	(l)
reactant	product	(g)	(aq)

The fuel for the space shuttle is hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) **(9)** reactant, oxygen is a(n) **(10)** reactant, and water vapor is a(n) **(11)** product. In a chemical equation for this reaction, a(n) **(12)** arrow is used to separate hydrogen and oxygen from water vapor and energy. A(n) **(13)** plus sign is used to separate the symbols for hydrogen and oxygen. A(n) **(14)** (g) symbol is used to tell the state of hydrogen in the reaction, a(n) **(15)** (g) symbol is used for the state of oxygen, and a(n) **(16)** (g) symbol is used for the state of water vapor.

Chemical Reactions Study Guide Answers

**Barron's Educational Series, Mark
Kernion, Joseph A. Mascetta**



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Chemistry Equations And Answers (Speedy Study Guides) Speedy Publishing, 2015-04-24 Chemistry is a difficult subject to fully comprehend with its equations and scientific laws Trying to digest an entire book in one semester is a tough job but with the help of study guides like these you can absorb information in chemistry much more effectively This guide covers chemical equations including examples potential problems and solutions

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Study Guide for Introductory Chemistry Iris Stovall,1990

Organic Chemistry Michael B. Smith,2016-03-09 Based on the premise that many if not most reactions in organic chemistry can be explained by variations of fundamental acid base concepts Organic Chemistry An Acid Base Approach provides a framework for understanding the subject that goes beyond mere memorization Using several techniques to develop a relational understanding it helps students fully grasp the essential concepts at the root of organic chemistry This new edition was

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