



1-27 A column was used to strip low-volatile materials from a high temperature heat transfer fluid. During a maintenance procedure, water was trapped between two valves. During normal operation, one valve was opened and the hot oil came in contact with the cold water. The result was almost sudden vaporization of the water, followed by considerable damage to the column.

Consider liquid water at 25°C and 1 atm. How many times does the volume increase if the water is vaporized and vapor at 100°C and 1 atm?

Solution:

From the steam tables at 25°C and 1 atm, $v_f = 1.000 \text{ cm}^3/\text{gm}$. At 100°C and 1 atm, $v_g = 1673.0 \text{ cm}^3/\text{gm}$. Thus, the volume expansion ratio between the liquid and the vapor is 1673 times. Liquid water exposed to hot oil will flash explosively, leading to considerable damage due to the expansion of the liquid.

1-28

$$\begin{aligned}
 \text{(a) } A_{\text{area}} &= \text{of tank roof} = \pi \frac{r^2}{4} = \frac{(3.14)(30)^2}{4} \\
 &= 706.5 \text{ ft}^2 \\
 &= (706.5)(144) = 1.017 \times 10^5 \text{ in}^2
 \end{aligned}$$

$$P = F/A = \frac{200 \text{ lb}}{1.017 \times 10^5 \text{ in}^2} = 1.97 \times 10^{-3} \text{ psia}$$

$$\begin{aligned}
 \text{Since } 1 \text{ atm} &= 14.7 \text{ psia} = 33.9 \text{ ft of H}_2\text{O} \\
 &= 406.9 \text{ in of water}
 \end{aligned}$$

Therefore

$$P = (1.97 \times 10^{-3} \text{ psia}) \left(\frac{406.9 \text{ in H}_2\text{O}}{14.7 \text{ psia}} \right)$$

$$P = 0.054 \text{ in. of water (Not much)}$$

Chemical Process Safety 3rd Edition Solution Manual

Zhe Xu



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