



CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons .

General Equilibrium		
<p>General Equilibrium:</p> <p>Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol \rightleftharpoons.</p> <p>Equilibrium Constant (K_c)</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Equilibrium Constant (K_c)</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$	<p>Graphical Representation of Equilibrium</p>
Equilibrium Constant, K_c		
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Equilibrium Constant, K_p		
<p>Equilibrium Constant, K_p</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	<p>Equilibrium Constant, K_p</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$	<p>Equilibrium Constant, K_p</p> <p>For a general reaction:</p> $aA + bB \rightleftharpoons cC + dD$ <p>The equilibrium constant is given by:</p> $K_p = \frac{P_C^c P_D^d}{P_A^a P_B^b}$
Le Chatelier's Principle: the effect of changing conditions on equilibrium		
<p>Concentration:</p> <ul style="list-style-type: none"> Increasing the concentration of a reactant shifts the equilibrium to the right. Increasing the concentration of a product shifts the equilibrium to the left. 	<p>Pressure:</p> <ul style="list-style-type: none"> Increasing the pressure shifts the equilibrium to the side with fewer moles of gas. Decreasing the pressure shifts the equilibrium to the side with more moles of gas. 	<p>Temperature:</p> <ul style="list-style-type: none"> Increasing the temperature shifts the equilibrium to the endothermic direction. Decreasing the temperature shifts the equilibrium to the exothermic direction.

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