

Pre-Lab Assignment

Kinetics of the Reaction of Folate Reductase with Fluoride

1. Write down the general expression for the dependence of a reaction's rate on the concentration of the reacting species.

Remember to include, wherever it is appropriate, clearly stated assumptions that provide a basis for the following generalizations. The following is a guide to help you, but it is not a substitute for the discussion in your textbook. It is your responsibility to ensure that you understand the following concepts and their relationships.

2. Write down the rate law for a reaction that is first order in A and second order in B. For a given reaction, write down the general rate law for a reaction that is first order in A and second order in B.

$$\text{Rate} = k[A]^1[B]^2 = k[A][B]^2$$

$$\text{Rate} = k[A][B]^2$$

3. Suppose you know that the reaction is first order in A and second order in B. Write down the general rate law for a reaction that is first order in A and second order in B. Write down the general rate law for a reaction that is first order in A and second order in B.

$$\text{Rate} = k[A]^1[B]^2 = k[A][B]^2$$

$$\text{Rate} = k[A][B]^2$$

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Ian M. Campbell



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